

Document downloaded from:

<http://hdl.handle.net/10251/189343>

This paper must be cited as:

López-Maya, E.; Padial, NM.; Castells-Gil, J.; Ganivet, CR.; Rubio-Gaspar, A.; García Cirujano, F.; Almora-Barrios, N.... (2021). Selective Implantation of Diamines for Cooperative Catalysis in Isorectular Heterometallic Titanium-Organic Frameworks. *Angewandte Chemie International Edition*. 60(21):11868-11873. <https://doi.org/10.1002/anie.202100176>



The final publication is available at

<https://doi.org/10.1002/anie.202100176>

Copyright John Wiley & Sons

Additional Information

This is the peer reviewed version of the following article: Selective Implantation of Diamines for Cooperative Catalysis in Isorectular Heterometallic Titanium-Organic Frameworks, which has been published in final form at <https://doi.org/10.1002/anie.202100176>. This article may be used for non-commercial purposes in accordance with Wiley Terms and Conditions for Self-Archiving.

Selective implantation of diamines for cooperative catalysis in isorecticular heterometallic titanium-organic frameworks

Elena López-Maya,^[a] Natalia M. Padial,^[a] Javier Castells-Gil,^[a] Carolina R. Ganivet,^[a] Francisco G. Cirujano,^[a] Neyvis Almora-Barrios,^[a] Sergio Tatay,^[a] Sergio Navalón^[b] and Carlos Martí-Gastaldo*^[a]

We introduce the first example of isorecticular titanium-organic frameworks, MUV-10 and MUV-12, to show how the different affinity of hard Ti(IV) and soft Ca(II) metal sites can be used to direct selective grafting of amines to the last. This enables the combination of Lewis acid titanium centers and -NH₂ Brønsted sites in two sizeable pores for cooperative cycloaddition of CO₂ to epoxides at room temperature and atmospheric pressure. The selective grafting of molecules to heterometallic clusters adds up to the pool of methodologies available for controlling the positioning and distribution of chemical functions in precise positions of the framework required for definitive control of pore chemistry.

Metal–Organic Frameworks (MOFs) have arisen as promising materials for cooperative catalysis due to the combination of high surface areas,^[1] dense number of open metal sites (OMS),^[2] and crystalline structures for controllable arrangements of active sites and additional catalytic units.^[3] Cooperative catalysis relies on the acceleration of the reaction rate by simultaneous activation of two or more reactants by different catalytic sites. These processes are a common feature in enzymes, that have naturally evolved to make use of the active center and the surrounding environment for tuning their catalytic activity.^[4,5] Cooperativity in MOFs can be achieved from the interplay of open metal sites with organic co-catalysts, that can be incorporated to the solid either by using functionalized linkers during the synthesis^[6–8] or loaded post-synthetically as functional guests ((**Figure 1a**)).^[9,10] However, these strategies sometime offer poor control over the distribution of active sites in the solid or weak host-guest interactions detrimental to activity or recyclability. Chemical design can be used to circumvent these shortcomings and produce cooperative catalysts that feature a tailorable distribution of active sites by direct implantation in precise positions of the framework.

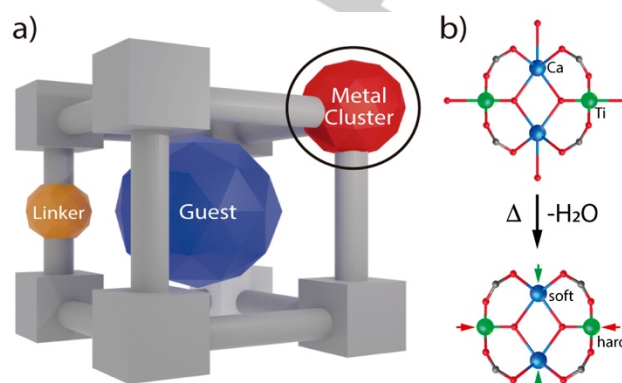


Figure 1. a) Different strategies to combined cooperative active sites in MOFs. b) Creation of two types of open metal sites (hard and soft) generated by thermal treatment.

water.^[11] Although the challenging solution chemistry of titanium has limited the isolation of crystalline frameworks, the number of Ti-MOFs reported has increased significantly in the last years.^[12–16] We recently reported the heterometallic MUV-10 framework [Ti^{IV}₃Ca^{II}₃(μ₃-O)₃(btc)₄(H₂O)₆] (btc = benzene-1,3,5-tricarboxylate).^[14] This material combines excellent thermal and chemical stabilities with a surface area close to 1000 m²·g⁻¹, all relevant properties from a catalytic standpoint. In addition, it features tetranuclear heterometallic Ti₂Ca₂ clusters with axially coordinated water molecules that enable straightforward generation of two types of open metal sites (OMS) by thermal treatment.

We demonstrate how the different chemistry in the affinity of unsaturated hard Ti(IV) and soft Ca(II) Lewis acid sites to coordinate amines can be used for selective implantation of basic Brønsted sites to the last (**Figure 1b**). Compared to other routes based on encapsulation or impregnation, the incorporation of the co-catalyst is here controlled by the chemistry of the metal-oxo cluster and is thus compatible with the isorecticular expansion of the framework. To exemplify this advantage, we extend this method to MUV-12, a new material built from the interlinking of the same Ti₂Ca₂ with an extended tricarboxylic linker, that permits obtaining enlarged pores to generate a cooperative catalyst for quantitative cycloaddition of CO₂ to epoxides at room temperature and atmospheric pressure that prevents the use of an additional molecular co-catalyst.

MUV-10 was synthesized according to the reported method.^[14] Phase purity was confirmed by chemical analysis and powder X-Ray diffraction (PXRD). We used water adsorption isotherms and FT-IR spectra of CO adsorption to confirm the creation of OMS by thermal treatment in vacuum (**Figure S1 and S2**). MUV-10 displays sodalite-type cages of 1.2 nm in diameter intersected in the 3D by channels with 0.5 nm pore windows (**Figure 2a**), only compatible with the grafting of small diamines. Accordingly, we reacted desolvated MUV-10 crystals with 1,3-diaminopropane

[a] Dr E. López-Maya, Dr N. M. Padial, Dr J. Castells-Gil, Dr C. R. Ganivet, Dr F. G. Cirujano, Dr. N. Almora-Barrios, Dr. S. Tatay and Dr C. Martí-Gastaldo

Functional Inorganic Materials team, Instituto de Ciencia Molecular, Universidad de Valencia. Catedrático Jose Beltrán-2, 46980, Paterna (Spain); e-mail: carlos.marti@uv.es

[b] Dr. S. Navalón

Departamento de Química e Instituto Universitario de Tecnología Química CSIC-UPV, Universitat Politècnica de València, c/ Camino de Vera, 46022, Valencia (Spain).

Supporting information for this article is given via a link at the end of the document.

Here we have chosen a titanium organic framework to demonstrate this concept. Despite MOFs often suffer from poor chemical stability that limits their use as heterogeneous catalysts the incorporation of hard, oxophilic Ti or Zr(IV) ions to their structure provides stronger bonds for exceptional stability in

(DAP) in anhydrous toluene for 12 hours at room temperature under inert atmosphere to produce MUV-10-DAP. This method is very similar to that used for appending diamines to the metal centres (Mg, Mn, Zn) in MOF-74.^[17,18] The PXRD of MUV-10-DAP confirms that the appending of DAP is respectful with the structure of the framework (**Figure 2b**). The incorporation of the diamine in the solid is consistent with the presence of characteristic C-H and N-H amine vibrations at 2864 and 3280 respectively cm^{-1} in the FT-IR spectrum (**Supplementary Section S2.2**). The extent of DAP loading was evaluated with thermogravimetric analysis (TGA) and proton nuclear magnetic resonance $^1\text{H-NMR}$ of the digested solid. Both suggest the grafting of approximately 13 wt% of DAP for the unit formula $[\text{Ti}^{\text{IV}}_3\text{Ca}^{\text{II}}_3(\mu_3\text{-O})_3(\text{btc})_4(\text{DAP})_3]$, consistent with the CHN analysis. Compared to homometallic $\text{Mg}_2(\text{dobpdc})$ that results in complete metal grafting for M:DAP ratios of 1:1,^[19] our experimental 2:1 ratio suggests that only 50% of the metal sites are grafted. This agrees well with a selective implantation to the soft Ca(II) sites in the heterometallic Ti_2Ca_2 cluster, more prone to interact with the primary amine of DAP than the hard oxophilic Ti(IV) metals. We next analyzed the changes in the porosity of the solid with N_2 and CO_2 adsorption experiments. Compared to MUV-10, the surface area of MUV-10-DAP is reduced to close the 50% for a BET of $520 \text{ m}^2 \text{ g}^{-1}$ (**Figure 2c**). This drop coincides with the reduction in free pore volume calculated with Materials Studio for a structural model of MUV-10 with DAP molecules appended only to Ca(II) sites (**Table S1**). CO_2 isotherms also reveal an important decrease in gas uptake of close to 75% compared to MUV-10, but display a higher uptake in the low pressure region ($P < 100 \text{ Torr}$) indicative of a stronger interaction with CO_2 (**Figure S6**).^[20,21] This is confirmed by the isosteric heat of CO_2 adsorption of diamine appended MUV-10-DAP (61.3 kJ mol^{-1}) that almost triplicates the 27.7 kJ mol^{-1} of the pristine framework (**Figure 2d**). This experimental information suggested MUV-10-DAP might behave as a bifunctional heterogeneous catalyst from the combination of Lewis acid and Brønsted basic sites. Before testing its ability to activate the cycloaddition of CO_2 to epoxides in absence of a molecular co-catalyst, we deemed necessary to generate iodide ions that are known to activate ring opening.^[22] MUV-10-DAP crystals were suspended in dry CH_3CN and reacted with methyl iodide at 80°C for 48 h under N_2 atmosphere according to a reported method.^[22] We evaluated the incorporation of iodine by ICP analysis (**Supplementary Section S2.1**). Unfortunately, our results discarded a successful generation of iodine ions. This is likely due to the strict limitations imposed by the grafting of DAP to the pore openings and empty space available in MUV-10, which is quite small in origin (**Figure 2e** and **Table S1**).

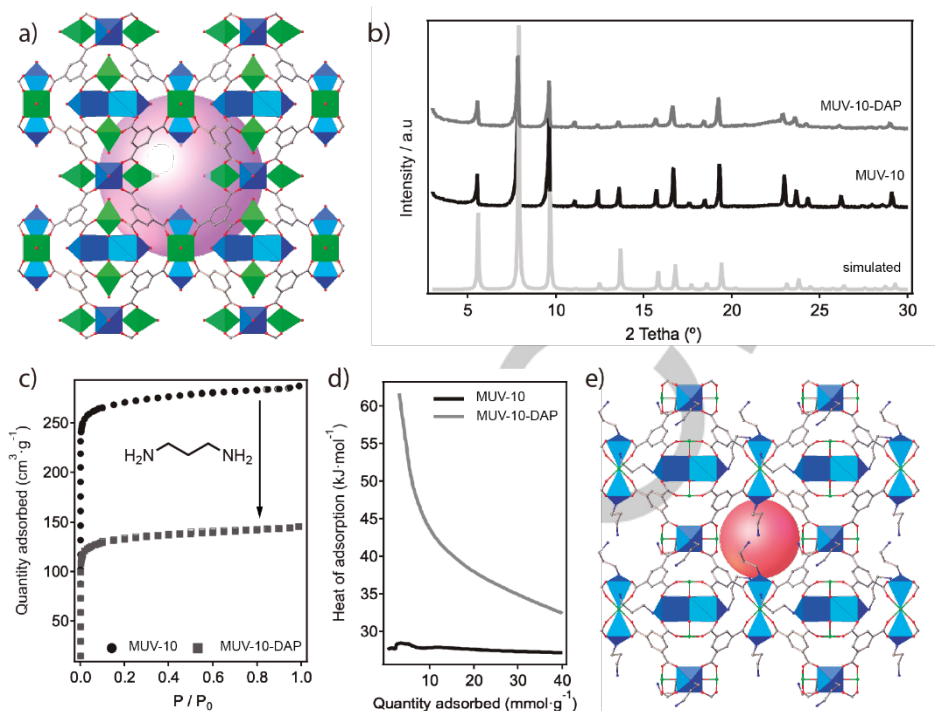


Figure 2. a) Structure of MUV-10 featuring sodalite-type octahedral cages of 1.2 nm in diameter (lavender sphere). b) PXRD patterns of MUV-10 and MUV-10-DAP after reaction with the diamine. c) N_2 isotherms showing the drop of surface area as result of the incorporation of DAP (MUV-10-DAP), which also induces a drastic increase in the d) CO_2 isosteric heat of adsorption. e) The simulated structure of MUV-10-DAP reveals a reduction in the pore windows and internal cavity (red sphere) of close to 50% that limits substrate uptake.

This limitation encouraged us to synthesize an isorecticular framework that combined superior porosity and heterometallic clusters amenable to selective DAP grafting. The reticulation of organic linkers of the same geometry and connectivity but variable size permits controlling the metrics and functionality of a targeted framework topology, that is in turn determined by the structural information encoded in the inorganic node or secondary building unit (SBU).^[23] Reticular chemistry has been incredibly fruitful in the design of multiple MOFs^[24] but the synthetic difficulties inherent to the chemistry of titanium in solution have neglected the discovery of persistent SBUs for the synthesis of isorecticular Ti-MOFs thus far. We synthesized 3,5-tris(4-carboxyphenyl)benzene (H_3btb) according to the reported protocol (**Figure 3a**, **Supplementary Section S3.1**). Reaction of btb with calcium chloride and titanium(IV) isopropoxide at 120°C in N,N -dimethylformamide (DMF) by using acetic acid as modulator yields colourless, octahedral crystals of ca. $25 \mu\text{m}$ in size (**Figure 3b**). Single-crystal XRD measurements revealed that MUV-12 $[\text{Ti}^{\text{IV}}_3\text{Ca}^{\text{II}}_3(\mu_3\text{-O})_3(\text{btb})_4(\text{H}_2\text{O})_6]$ is a doubly interpenetrated framework isorecticular to MUV-10. It crystallizes in the cubic space group $I m \bar{3}$ ($a = 26.4010 \text{ \AA}$) and, despite interpenetration, doubles the accessible void space with pore diameters of 2.4 nm (**Figure 3c**). Besides overcoming the spatial limitations intrinsic to MUV-10, we argued framework interpenetration might be beneficial from a catalytic standpoint due to a higher density of active sites and enhanced robustness.^[25]

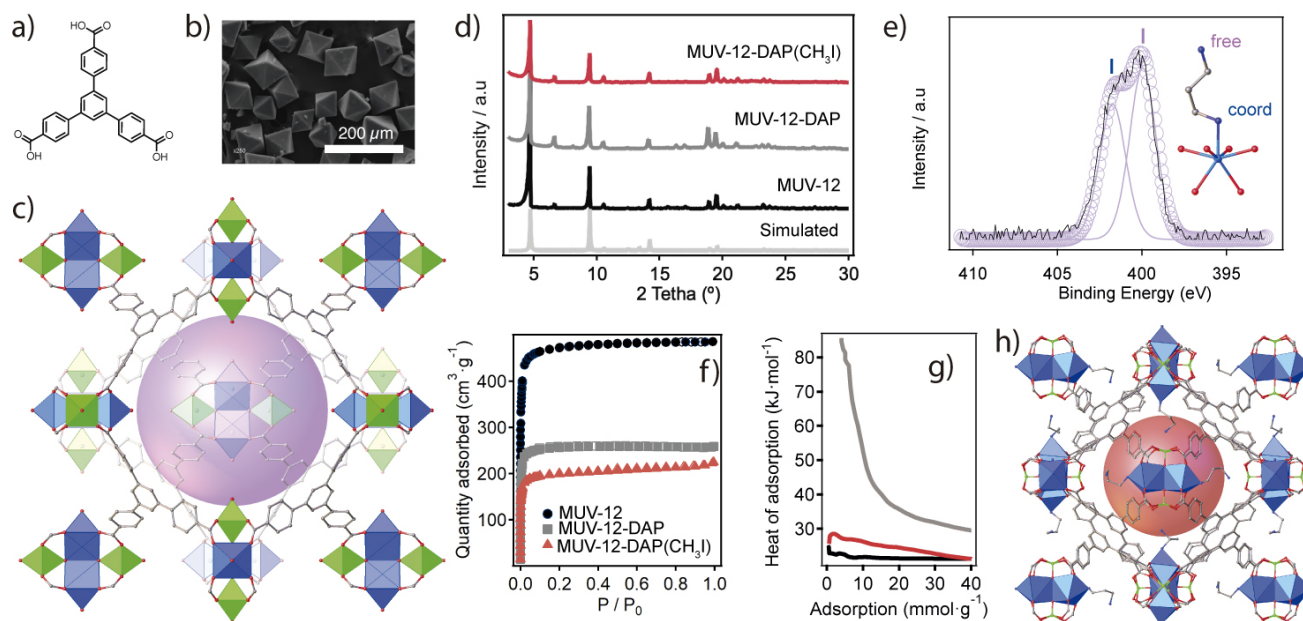


Figure 3. a) Structure of btb linker. b) SEM images of MUV-12 crystals. c) Structure of MUV-12 featuring an interpenetrated structure and sodalite-type octahedral cages of ca. 2.2 nm in diameter (lavender sphere). d) PXRD patterns of MUV-12, MUV-12-DAP and MUV-12-DAP(CH₃I). e) XPS of the N 1s region in MUV-12-DAP f) N₂ isotherms showing the drop of surface area as result of the incorporation of DAP (MUV-12-DAP) as well as the incorporation of alkyl groups and triiodide (MUV-12-DAP(CH₃I)) g) CO₂ isosteric heat of adsorption of MUV-12, MUV-12-DAP and MUV-12-DAP(CH₃I). h) Simulated pore structure of MUV-12-DAP reveals a reduction in the pore windows and internal cavity (red sphere) of close to 50% that limits substrate uptake.

We synthesized a fresh batch of MUV-12 and confirmed phase purity by PXRD and CHN (**Supplementary Section S3**), followed by DAP grafting and thorough washing according to the same method used for MUV-10. It is worth noticing that all materials used in this work were subjected to a thermal treatment to prevent the presence of uncoordinated diamine in the pore as confirmed by X-Ray Photoelectron Spectroscopy (XPS, **Supplementary Figure S11**). TGA and ¹H-NMR suggest the appending of approximately 7.8 weight percent (wt%) of DAP to the solid for a M:DAP ratio of 2:1. This value is equivalent to that obtained for MUV-10 and further supports the selective implantation to Ca(II) sites for the formation of [Ti^{IV}₃Ca^{II}₃(μ₃-O)₃(btb)₄(DAP)₃]. This unit formula was also corroborated by CHN analysis of the crystals. It is worth noting that MUV-12 does not suffer from the limitations to the diffusion and grafting of the diamine imposed by MUV-10. As result, it could incorporate more easily a stoichiometric M:DAP ratio unless the differences in affinity of both metals for diamine coordination are quite severe. As supported by MUV-12-DAP PXRD data, DAP grafting has a negligible impact on the structure of MUV-12 (**Figure 3d**). We used nitrogen 1s X-ray photoelectron spectroscopy (XPS) to confirm the appending of the diamine to the framework. **Figure 3e** shows the N1s high-resolution XPS spectrum of MUV-12-DAP. The N1s signal confirms the presence of nitrogen and reveals two different chemical environments in roughly 1:1 ratio. The contribution centered at 400.0 and 402.3 eV can be attributed to free primary -NH₂ and secondary amine groups bound to the metal, respectively.^[26,27] The effect of DAP grafting on the porosity of MUV-12 is quite similar to that observed for MUV-10. As shown in **Figure 3f**, the experimental BET surface area of 2100 m²·g⁻¹ in MUV-12 drops to around 50 % in MUV-12-DAP (980 m²·g⁻¹). The analysis of CO₂ isotherms at 273, 283 and 293 K reveals that MUV-12-DAP outperforms the non-aminated framework at low-P regimes, indicative of enhanced adsorption (**Supplementary Figure S14**). This is consistent with the drastic increase in the isosteric heat of CO₂ adsorption (Q_{st}) observed

after appending of DAP (**Figure 3g**). MUV-12-DAP Q_{st} of 84.6 kJ mol⁻¹, reaches close to four times the value of pristine MUV-12 (24.3 kJ mol⁻¹).

The simulated structural model of MUV-12-DAP reveals a free volume of 2176 Å³ with accessible pore diameters of close to 20 Å and pore windows of 7 Å (**Figure 3h**). Compared to the spatial limitations intrinsic to MUV-10-DAP, this value shall enable generation of nucleophilic iodide anions in the pore more easily. Accordingly, we reacted MUV-12-DAP crystals with CH₃I by following the same procedure described above for its isorecticular counterpart. The resulting material MUV-12-DAP(CH₃I) was analyzed with ICP (**Supplementary Section S3.1**) and energy dispersive X-ray (EDX) to confirm the generation of iodine ions. In addition, mapping images show the homogeneous distribution of iodine across the crystal and scanning electronic microscopy (SEM) images verify that the size and morphology of the crystals is maintained in MUV-12-DAP(CH₃I) discarding significant damage (**Figure Supplementary S16**). Nonetheless, CHN analysis reveal that this post-synthetic modification induces a loss of close to 60% of the nitrogen originally present in MUV-12-DAP. The PXRD of MUV-12-DAP(CH₃I) confirms that the framework can survive multistep PSM with negligible changes to the original diffraction pattern (**Figures 3d and S10**).

Surprisingly, the incorporation of triiodide ions does not lead to a high decrease in BET surface area value (850 m²·g⁻¹) for MUV-12-DAP (**Figure 3f**). This is due to the sublimation of iodine during the activation of the framework prior to gas adsorption, as confirmed by the EDX analysis of the material after the isotherm that reveals a complete loss of iodine (**Figure Supplementary S16e**). By contrast, the treatment with CH₃I reduced CO₂ uptake at 20 °C (40 cm³·g⁻¹). The weaker interaction with CO₂ denoted by the absence of steep increases in the gas uptake at low-P indicate the modification of -NH₂ groups (**Figure Supplementary S15**). This is further corroborated by the decrease of the heat of

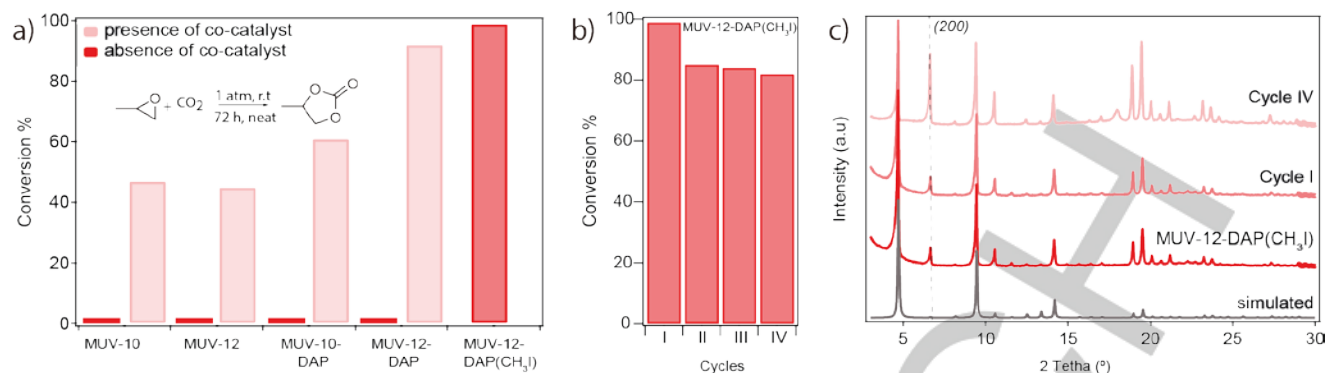


Figure 4. a) Cyclic carbonate conversion through cycloaddition of CO₂ into propylene oxide using MUV-10, MUV-12, MUV-10-DAP and MUV-12-DAP in presence (8% mol) and in absence of co-catalyst and MUV-12-DAP(CH₃I) only in absence of co-catalyst. b) Recycled experiments of cycloaddition of CO₂ with propylene oxide catalyzed by MUV-12-DAP(CH₃I). c) PXRD of MUV-12 simulated, MUV-12-DAP(CH₃I), MUV-12-DAP(CH₃I) after the first cycle and MUV-12-DAP(CH₃I) after fourth cycle of catalysis.

adsorption at zero coverage to 28.4 kJ mol⁻¹ (Figure 3g), slightly above pristine MUV-12.

To illustrate the value of diamine grafting and incorporation of halides to the framework, we tested the catalytic activity of the MUV-10 and MUV-12 families in the formation of cyclic carbonates by the cycloaddition of CO₂ to epoxides. There are previous reports on the use of MOFs to catalyze this reaction but they generally require high pressures, high temperatures or the use of a co-catalyst.^[22,28–30] This renders a complicated scenario that is not ideal from an economical and environmental standpoint. The design of more efficient catalysts capable of CO₂ conversions at room temperature and atmospheric pressure without a co-catalyst is highly desirable. To demonstrate the efficiency of our material, we tested the cycloaddition of CO₂ to form propylene carbonate at rt and 1 atm of pressure (Figure 4a). The conversion with pristine MUV-10 and MUV-12 is negligible in absence of a co-catalyst. Addition of tetrabutylammonium bromide (TBAB) is required to activate the reaction, reaching conversions slightly above 40% in both materials. We do not observe significant differences that could be attributed to the superior porosity of the extended MUV-12 framework. More clear differences can be observed for the changes in catalytic activity after the appending of DAP. Just like their non-aminated counterparts, MUV-10-DAP and MUV-12-DAP are inactive in absence of TBAB but display comparatively higher conversions after the addition of the co-catalyst. This boost is particularly significant for MUV-12-DAP, that reaches close to 90% in conversion compared to the 60% of MUV-10-DAP. We argue these changes between isostructural frameworks reveal the positive effect of -NH₂ Brønsted sites to activate CO₂ for the formation of the corresponding cyclic carbonate.^[31] This effect is more acute for MUV-12 possibly from its superior porosity area and interpenetrated structure. This combination results in a superior CO₂ uptake at 20 °C (46 cm³·g⁻¹) and a higher density of appended DAP molecules with the associated increase in Q_{st} for more efficient CO₂ activation. Finally, MUV-12-DAP(CH₃I) displays the highest efficiency of the series and is the only material that reaches complete conversion in absence of the molecular co-catalyst. It should be noted that in the control experiment with the physical mixture of MUV-12, DAP and CH₃I the conversion obtained was only 40%, thus suggesting that the implantation of DAP plays a key role for the subsequent generation of iodine ions inside the pore.

This drastic change in performance of MUV-12-DAP(CH₃I) can be only explained as result of the incorporation triiodide ions to the framework, thus preventing the use of an additional reagent.

MUV-12-DAP(CH₃I) behaves as a cooperative catalyst from the combination of Lewis acid Ti(IV) and nucleophilic I₃⁻ anions, both cooperating in the pores of the framework for synergetic interaction in a confined space. The established mechanism for this reaction involves a Lewis acid site to activate the epoxide towards the nucleophilic attack by the halide (Scheme S1). Whereas the density of Lewis sites is fixed by the heterometallic Ti₂Ca₂ cluster, our method permits producing MUV-12-DAP(CH₃I) catalysts with variable amounts of nucleophile to evaluate the effect of these triiodide anions on the overall performance. By changing the temperature and reaction time used for the synthesis of MUV-12-DAP(CH₃I) (Table S4), we synthesized materials with different amounts of I₃⁻. The catalytic tests confirm that epoxide conversion in our case is controlled by the nucleophile, in good agreement with the poor efficiencies encountered for MUV-10-DAP and MUV-12-DAP in absence of TBAB.

To put our results in a general context, we have summarized the catalytic activity reported for other MOFs used for the transformation of epoxides into cyclic carbonates at standard pressure (Table Supplementary S5). MUV-12-DAP(CH₃I) is the only one that displays quantitative conversion at room temperature, just comparable to the activity displayed by composites of mesoporous MIL-101 or PCN-333 prepared by infiltration with an ionic polymer.^[10] Compared to this approach that can induce the co-precipitation of the polymer at the surface of the crystals, our method enables direct implantation of the cooperating sites in the confined space by using the chemistry of the SBU. To confirm that the CO₂ transformation could be ascribed exclusively to in-pore reactivity, we also tested the same reaction under equivalent conditions with other substrates of increasing sizes (Table Supplementary S5). We observe a progressive decrease with the size of the substrate down to 30% in conversion for 1,2-epoxybutane, 7% for butyl glycidyl ether and below 3% for bigger substrates, thus ruling out surface reactivity. Also important is the absence of an additional component simplifies the recovery of the catalyst for additional tests. Figure 4b shows the cyclability of MUV-12-DAP(CH₃I), which maintains the conversion of propylene carbonate above 80% upon four consecutive cycles. According to the PXRD of the material after each test (Figure 4c) this small loss in activity is not indicative of the degradation of the framework. This is more likely due to the blocking of the pores by the product of the reaction in view of the changes in relative intensity of the diffraction peaks in accordance with different loadings of the MUV-12-DAP(CH₃I) pores.^[32]

We introduce here the first family of isorecticular titanium frameworks, MUV-10 and MUV-12, to demonstrate the ability of heterometallic Ti_2Ca_2 clusters to direct the selective implantation of diamines to the soft Ca(II) sites in the inorganic node. This results in the combination of Lewis acid Ti(IV) centers and $-NH_2$ Brønsted sites in two pores of increasing metrics for cooperative catalysis. Chemical modification of MUV-12 enables quantitative cycloaddition of CO_2 to propylene oxide at room temperature and atmospheric pressure for complete conversion in absence of additional co-catalysts and with excellent cyclability. This selective grafting of molecules to heterometallic clusters adds up to the pool of methodologies available for controlling the positioning and distribution of functional moieties in MOFs. We are confident The combined use of the empty space offered by these porous frameworks, alongside the different chemistries intrinsic to metal clusters and organic linkers, offers a promising playground for producing chemically complex environments by implantation of multiple components for unprecedented opportunities in synergistic phenomena.^[33]

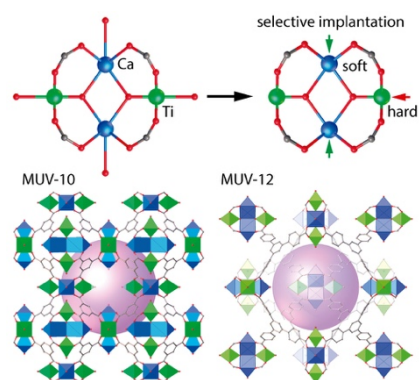
Acknowledgements

This work was supported by the EU (ERC Stg Chem-fs-MOF 714122) and Spanish government (CTQ2017-83486-P & MDM-2015-0538). E.L.-M. thanks the Spanish government for a Juan de la Cierva Fellowship (FJCI-2017-32956). N.M.P. thanks the European Union for a Marie Skłodowska-Curie Global Fellowship (H2020-MSCA-IF-2016-GF-749359-EnanSET). F. G. C. acknowledges the support of a fellowship from "la Caixa" Foundation (ID 100010434). The fellowship code is LCF/BQ/PI19/11690011.

Keywords: titanium-organic frameworks • carbon dioxide conversion • cooperative catalysis • isorecticular chemistry • heterometallic cluster

- [1] Z. Chen, P. Li, R. Anderson, X. Wang, X. Zhang, L. Robison, L. R. Redfern, S. Moribe, T. Islamoglu, D. A. Gómez-Gualdrón, et al., *Science* **2020**, *368*, 297–303.
- [2] Y. S. Wei, M. Zhang, R. Zou, Q. Xu, *Chem. Rev.* **2020**, DOI 10.1021/acs.chemrev.9b00757.
- [3] Y. B. Huang, J. Liang, X. S. Wang, R. Cao, *Chem. Soc. Rev.* **2017**, *46*, 126–157.
- [4] X. H. Matthew D. Wodrich, *Nat. Rev. Chem.* **2017**, *2*, 1–7.
- [5] B. Le Bailly, *Nat. Rev. Chem.* **2018**, *2*, 385–397.
- [6] P. Z. Li, X. J. Wang, J. Liu, J. S. Lim, R. Zou, Y. Zhao, *J. Am. Chem. Soc.* **2016**, *138*, 2142–2145.
- [7] C. Li, H. Tang, Y. Fang, Z. Xiao, K. Wang, X. Wu, H. Niu, C. Zhu, H. C. Zhou, *Inorg. Chem.* **2018**, *57*, 13912–13919.
- [8] B. Li, K. Leng, Y. Zhang, J. J. Dynes, J. Wang, Y. Hu, D. Ma, Z. Shi, L. Zhu, D. Zhang, et al., *J. Am. Chem. Soc.* **2015**, *137*, 4243–4248.
- [9] X. Li, B. Zhang, L. Tang, T. W. Goh, S. Qi, A. Volkov, Y. Pei, Z. Qi, C. K. Tsung, L. Stanley, et al., *Angew. Chemie - Int. Ed.* **2017**, *56*, 16371–16375.
- [10] B. Aguilá, Q. Sun, X. Wang, E. O'Rourke, A. M. Al-Enizi, A. Nafady, S. Ma, *Angew. Chem Int. Ed.* **2018**, *57*, 10107–10111.
- [11] A. J. Howarth, Y. Liu, P. Li, Z. Li, T. C. Wang, J. T. Hupp, O. K. Farha, *Nat. Rev. Mater.* **2016**, *1*, 1–15.
- [12] N. M. Padiál, B. Lerma-Berlanga, N. Almora-Barrios, J. Castells-Gil, I. da Silva, M. de la Mata, S. I. Molina, J. Hernández-Saz, A. E. Platero-Prats, S. Tatay, et al., *J. Am. Chem. Soc.* **2020**, *142*, 6638–6648.
- [13] J. Castells-Gil, N. M. Padiál, N. Almora-Barrios, I. Da Silva, D. Mateo, J. Albero, H. García, C. Martí-Gastaldo, *Chem. Sci.* **2019**, *10*, 4313–4321.
- [14] J. Castells-Gil, N. M. Padiál, N. Almora-Barrios, J. Albero, A. R. Ruiz-Salvador, J. González-Platas, H. García, C. Martí-Gastaldo, *Angew. Chem Int. Ed.* **2018**, *57*, 8453–8457.
- [15] L. Zou, D. Feng, T. F. Liu, Y. P. Chen, S. Yuan, K. Wang, X. Wang, S. Fordham, H. C. Zhou, *Chem. Sci.* **2016**, *7*, 1063–1069.
- [16] H. L. Nguyen, F. Gándara, H. Furukawa, T. L. H. Doan, K. E. Cordova, O. M. Yaghi, *J. Am. Chem. Soc.* **2016**, *138*, 4330–4333.
- [17] P. J. Milner, R. L. Siegelman, A. C. Forse, M. I. Gonzalez, T. Runčevski, J. D. Martell, J. A. Reimer, J. R. Long, *J. Am. Chem. Soc.* **2017**, *139*, 13541–13553.
- [18] C. M. McGuirk, R. L. Siegelman, W. S. Drisdell, T. Runčevski, P. J. Milner, J. Oktawiec, L. F. Wan, G. M. Su, H. Z. H. Jiang, D. A. Reed, et al., *Nat. Commun.* **2018**, *9*, 1–10.
- [19] P. J. Milner, R. L. Siegelman, A. C. Forse, M. I. Gonzalez, T. Runčevski, J. D. Martell, J. A. Reimer, J. R. Long, *J. Am. Chem. Soc.* **2017**, *139*, 13541–13553.
- [20] C. Montoro, E. García, S. Calero, M. A. Pérez-Fernández, A. L. López, E. Barea, J. A. R. Navarro, *J. Mater. Chem.* **2012**, *22*, 10155–10158.
- [21] A. Demessence, D. M. D'Alessandro, M. L. Foo, J. R. Long, *J. Am. Chem. Soc.* **2009**, *131*, 8784–8786.
- [22] J. Liang, R. P. Chen, X. Y. Wang, T. T. Liu, X. S. Wang, Y. B. Huang, R. Cao, *Chem. Sci.* **2017**, *8*, 1570–1575.
- [23] M. J. Kalmutzki, N. Hanikel, O. M. Yaghi, *Sci. Adv.* **2018**, *4*, 1–16.
- [24] M. Eddaoudi, J. Kim, N. Rosi, D. Vodak, J. Wachter, M. O'Keeffe, O. M. Yaghi, *Science (80-)* **2002**, *295*, 469–472.
- [25] Z. Lin, Z. M. Zhang, Y. S. Chen, W. Lin, *Angew. Chemie - Int. Ed.* **2016**, *55*, 13739–13743.
- [26] W. Cao, Z. Wang, Q. Zeng, C. Shen, *Appl. Surf. Sci.* **2016**, *389*, 404–410.
- [27] J. Ederer, P. Janoš, P. Ecorchard, J. Tolasz, V. Štengl, H. Beneš, M. Perchacz, O. Pop-Georgievski, *RSC Adv.* **2017**, *7*, 12464–12473.
- [28] J. Zhu, P. M. Usov, W. Xu, P. J. Celis-Salazar, S. Lin, M. C. Kessinger, C. Landaverde-Alvarado, M. Cai, A. M. May, C. Slebodnick, et al., *J. Am. Chem. Soc.* **2018**, *140*, 993–1003.
- [29] H. He, Q. Sun, W. Gao, J. A. Perman, F. Sun, G. Zhu, B. Aguilá, K. Forrest, B. Space, S. Ma, *Angew. Chemie - Int. Ed.* **2018**, *57*, 4657–4662.
- [30] C. Huang, J. Dong, W. Sun, Z. Xue, J. Ma, L. Zheng, C. Liu, X. Li, K. Zhou, X. Qiao, et al., *Nat. Commun.* **2019**, *10*, 1–10.
- [31] Q. Han, B. Qi, W. Ren, C. He, J. Niu, C. Duan, *Nat. Commun.* **2015**, *6*, 1–8.
- [32] H. S. Cho, J. Yang, X. Gong, Y. B. Zhang, K. Momma, B. M. Weckhuysen, H. Deng, J. K. Kang, O. M. Yaghi, O. Terasaki, *Nat. Chem.* **2019**, *11*, 562–570.
- [33] J. Castells-Gil, N. M. Padiál, N. Almora-Barrios, R. Gil-San-Millán, M. Romero-Ángel, V. Torres, I. da Silva, B. C. J. Vieira, J. C. Waerenborgh, J. Jagiello, J. A. R. Navarro, S. Tatay, C. Martí-Gastaldo, *Chem* **2020**, *6*, 3118–3131.

Entry for the Table of Contents



Selective grafting in heterometallic MOFs: We use the different affinity of Ti(IV) and Ca(II) sites to direct selective coordination of amines to the last. This enables the combination of Lewis acid and Brønsted sites in sizeable pores for cooperative cycloaddition of CO₂ to epoxides. MUV-12 displays full conversion of propylene oxide at room temperature and atmospheric pressure in absence of additional co-catalysts and with excellent cyclability.

Institute and/or researcher Twitter usernames: @FuniMAT, @ICMol_UV