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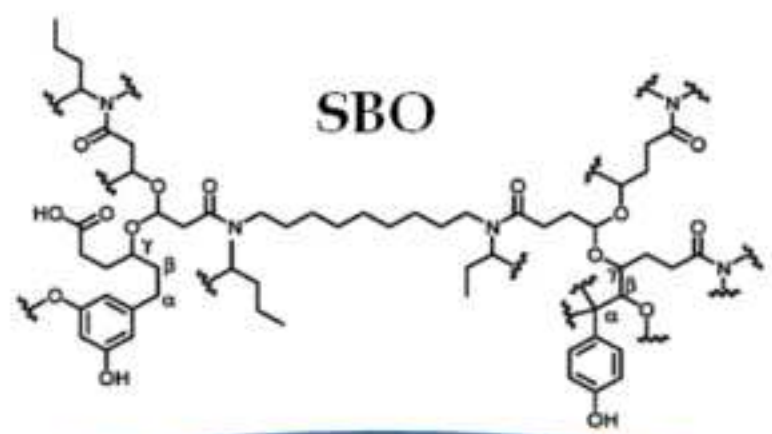
## Highlights

The use of SBO isolated from urban wastes for water treatment has been explored.

SBO are able to catalyze photobleaching of crystal violet

A CV-SBO complex might play a major role in the process

SBO favoured neutral photo-Fenton processes



**Crystal Violet**

## **Application of soluble bio-organic substances (SBO) as photocatalysts for wastewater treatment: sensitizing effect and photo-Fenton-like process**

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### **Abstract**

Soluble bioorganic surfactants (SBO) isolated from urban biorefuses have been investigated in photocatalytic processes for wastewater treatment using a cationic dye, namely crystal violet (CV) as target compound. The SBO have been found to enhance the photobleaching of CV solutions with an optimal SBO concentration of ca. 20 mg L<sup>-1</sup>. Mechanistic investigation based on chemical probes and changes in the absorption spectrum of CV in the presence of SBO seems to indicate that a complex formed between sensitizer and substrate plays a major role in the process. SBO have also been tested for the implementation of photo-Fenton processes at circumneutral media: The observed pseudo first order rate constants for CV decoloration was higher for UV/SBO/Fe(II)/H<sub>2</sub>O<sub>2</sub> ( $5.5 \times 10^{-3} \text{ min}^{-1}$ ) than for UV/Fe(II)/H<sub>2</sub>O<sub>2</sub> ( $1.5 \times 10^{-3} \text{ min}^{-1}$ ). A modified mechanism, in which HO· are less relevant than at the acidic medium, is in agreement with results of experiments carried out with chemical probes.

**Keywords:** cationic dyes; water treatment; photo-Fenton; organic biosurfactant; waste valorization

## **Introduction**

Dissolved organic matter (DOM) has received much attention from researchers in recent years because of its photochemical properties. In particular, humic and fulvic acids have been shown to generate, upon solar irradiation, oxidizing species such as hydroxyl radical [1, 2], singlet oxygen [3], superoxide radical ion [4] or reactive triplet excited states [5, 6]. These species can react with chemicals present in surface waters and participate in the most important abiotic processes for xenobiotics removal [7]. A review on the use of organic compounds with photochemical activity (namely dyes) for water treatment has been very recently published [8]. Unfortunately, DOM cannot be considered at commercial level as viable agent for the detoxification of wastewater due to the low concentration in water and soil.

Recent work reports how urban bio-wastes (UBW) have become a sustainable source of soluble bioorganic substances (SBO) having similar origin, chemical nature and photosensitizing properties as DOM [9]. It is interesting to note, that bio-wastes may have a potential beneficial fall out for environment. The potential use of SBO as photocatalyst for water detoxification of aqueous effluents may be considered a green process since it valorizes solid waste as a material of technological application.

The SBO are described as mixtures of macromolecules with weighted average molecular weight ( $M_w$ ) ranging from 67 to 463 kg mol<sup>-1</sup> and polydispersity indexes ( $M_w/M_n$ ) in the 6 to 53 range. When analyzing chemical composition data these macromolecules appear to be formed by long aliphatic C chains substituted by aromatic rings and several functional groups as COOH, CON, C=O, PhOH, O-alkyl, OAr, OCO,

OMe, and NRR', with R and R' being alkyl substituents. These organic moieties are the likely memory of the main constituents of the sourcing bio-organic refuse matter which are not completely mineralized during aging under aerobic fermentation conditions. For this reason, SBO may be considered to be the pristine material of DOM formed under longer aging conditions. The former, together with easier and continuous availability, seem to have some other interesting features. They can be obtained over a wide range of chemical composition depending on the type, location and treatment of the sourcing UBW [10]. Due to the presence of functional groups with strong metals' chelating power, the SBO contain a considerable fraction of minerals which are hard to separate from the organic matter. In principle some of these minerals, such as iron, could contribute to determine the SBO photochemical activity. Therefore, for their capacity to hold iron ions in solution at neutral-alkaline pH, SBO offer intriguing scope to investigate possible photo-Fenton-like processes occurring under pH conditions where iron ions are normally not soluble.

To fully understand the rationale motivating the present work, it should be considered that the Fenton reagent consists of a mixture of iron salts and hydrogen peroxide which is able to generate highly oxidizing species, such as hydroxyl radical [11]. The process is highly enhanced by irradiation and sunlight can be employed for this purpose with ecological and economical advantages [12]. Although this method has proven to be efficient, it is limited by the acidic pH that is required, partly due to the low solubility of iron at neutral media. In order to overcome this inconvenience, complexing agents such as humic acids have been employed [13-16]. Although previous papers have reported interesting photosensitizing behavior [17, 18], the applicability of SBO for photo-Fenton remains unexplored. The present paper is meant to contribute further insight into the applicability of SBO in photochemical processes

for wastewater treatment. For this purpose, Crystal Violet (CV), a cationic dye, was used as model compound. The role of SBO in the photobleaching of CV solutions was then studied in two different directions, namely as photosensitizer and as complexing agent to drive photo-Fenton processes under non acidic conditions.

## **2. Experimental**

### **2.1 Reagents**

Crystal Violet (Anedra),  $\text{Fe}(\text{ClO}_4)_3 \cdot n\text{H}_2\text{O}$  (Aldrich),  $\text{H}_2\text{O}_2$  (Perhydrol 30%, Merck), furfuryl alcohol (99%, Aldrich) and 2-propanol (Anedra) were analytical grade and used as received. Water of Milli-Q quality was used in all the experiments.

### **2.2 SBO isolation and characterization**

The investigated SBO, namely CVT230 was sourced from UBW sampled from the process lines of ACEA Pinerolese waste treatment plant in Pinerolo (Italy). The UBW was a material obtained in the compost production section from urban vegetable residues aged for 230 days. The UBW was processed further in a pilot plant made available from Studio Chiono e Associati in Rivarolo Canavese, Italy. This comprised an electrically heated mechanically stirred 500 L reactor, a 102 cm long x 10.1 cm diameter polysulfone ultrafiltration (UF) membrane with 5 kDa molecular weight cut-off supplied by Idea Engineering s.r.l., and a forced ventilation drying oven. UBW were digested 4 h at 60 °C, pH 13 and 4 V/w water/solid ratio. The liquid/solid mixture was allowed to settle to yield the supernatant liquid phase containing the soluble UBW hydrolyzate. The recovered liquid phase was circulated at 40 L h<sup>-1</sup> flow rate through the UF membrane operating with tangential flow at 7 bar inlet and 4.5 bar outlet pressure to yield a retentate with 5-10 % dry matter content. The concentrated retentate was finally

dried at 60 °C. The solid SBO product obtained in 15-30 % w/w yield, relatively to the starting UBW dry matter, was found to contain 72.1% (w/w) of volatile solids. The carbon content was 38.3 % which, in turn, could be classified according to the functional: aliphatic groups (37%), amines (7%), alkoxy (14%), anomeric (4%), aromatic (13%), phenolic (5%), phenoxy (2%), carboxylic (12%), amide (1%) and ketone (5%). It is remarkable that ca. 0.77% (w/w) of iron was found in the chemical composition of CVT230.

### **2.3 Experimental conditions**

Photochemical experiments were carried out in a 500 mL glass reactor equipped with a medium-pressure mercury arc lamp (Philips HPK 125 W) with a glass jacket that cut-off the radiation with wavelengths shorter than 300 nm. The solution was continuously bubbled with analytical air and the temperature of the system was kept at 25°C. The incident photon rate ( $P_0$ ), measured using potassium ferrioxalate as actinometer [19], was  $3.2 \times 10^{-6}$  einsteins  $s^{-1} L^{-1}$ . In all the experiments, the pH of the medium was adjusted to 7.0. The initial concentration of CV was 10 mg  $L^{-1}$  and the initial concentrations of SBO varied over the range from 1 to 200 mg  $L^{-1}$ . Samples were eventually taken from the reaction and their absorption spectra were recorded on a double-beam Shimadzu spectrophotometer using quartz cells of 1.0 cm optical path length.

### **2.4 Mathematical analysis**

Multivariate curve resolution-alternating least squares (MCR-ALS) was applied for investigating CV-SBO interactions. This procedure has already been successfully applied to the study of equilibria, when neither models nor prior information about the number and identity of the species involved were available [20, 21]. This method is



based on the resolution of the experimental data matrix,  $\mathbf{A}(m \times w)$ , into matrices  $\mathbf{C}(m \times n)$  and  $\mathbf{S}^T(n \times w)$ , by the iterative application of the following matrix product:  $\mathbf{A} = \mathbf{CS}^T + \mathbf{E}$ , where  $\mathbf{C}(m \times n)$  is the concentration profile matrix;  $\mathbf{S}^T(n \times w)$  is that containing the pure spectra of the chemical species involved, and  $\mathbf{E}(m \times w)$  represents the error matrix. Symbols  $m$ ,  $n$ , and  $w$  denote the BOS concentration for each spectrum, the number of absorbing species, and the recorded wavelengths, respectively. The value of  $n$  is usually unknown and, thus, singular value decomposition (SVD) was applied for the estimation of  $n$  [21].

### 3. Results and discussion

#### 3.1. Photolysis of CV in the presence of SBO

The role of SBO on the decoloration of CV in aqueous solution was investigated by recording the absorption spectra of CV at different irradiation times under the following experimental conditions:  $[\text{CV}] = 10 \text{ mg L}^{-1}$  and  $[\text{SBO}] = 25 \text{ mg L}^{-1}$  at  $\text{pH} = 7.0$ . Figure 1 shows a noticeable decrease in the band with a maximum at 532 nm, which can be associated to the removal of CV. The discoloration kinetics, based on the absorbance recorded at 532 nm, can be fitted to a pseudo-first order law (Figure 1, inset). Although this simple model cannot be considered to account in detail for the complexity involved in this kind of photoreactions, it can represent a useful tool for comparison purposes, allowing calculating an apparent bleaching constant ( $k_{\text{app}}$ ) for each experimental condition. In order to ascertain the possible role of direct photolytic degradation, the irradiation of CV solution was also performed in the absence of SBO. The trends reported in the inset in Fig.1, together with the calculated  $k_{\text{app}}$  of  $3.5 \times 10^{-3}$

$\text{min}^{-1}$  and  $5.8 \times 10^{-4} \text{ min}^{-1}$  in the presence and in the absence of SBO, respectively, clearly show that SBO promotes the indirect photochemical degradation of CV.

The effect of the concentration of SBO on the decoloration rate was investigated. Figure 2 shows the plot of  $k_{\text{app}}$  as a function of the concentration of added SBO. The bleaching rate of CV was observed to increase upon increasing the concentration of SBO, reaching a maximum value at  $[\text{SBO}] = 15 - 25 \text{ mg L}^{-1}$ . Beyond this point, further SBO additions lead to lower  $k_{\text{app}}$  values. Analogous behaviors have also been reported for heterogeneous photochemical systems, such as  $\text{TiO}_2$  [12] and for heterogenized organic photocatalysts [22]. In those cases, the existence of a maximum was attributed to a light screening effect at higher concentrations of the photocatalyst. However, in the system examined in this work other concentration-dependent processes cannot be ruled out, among them scavenging by SBO of the generated reactive species, or the formation of photoactive complexes between SBO and CV.

A series of experiments was therefore devoted to assess the role of the different reactive species formed by SBO upon irradiation. Degradation experiments were performed by irradiating solutions containing CV and SBO in the presence of either 2-propanol, an  $\text{HO}\cdot$  scavenger [23] or furfuryl alcohol, a singlet oxygen trapping agent [24]. None of these probes had any effect on the degradation rate (Figure 3). This implies that none of these reactive species plays a major role in the CV photobleaching process. On the other hand, when nitrogen purged (deoxygenated) solutions were employed, a faster CV decolorization was observed (Figure 3). Molecular oxygen is known to be a triplet quencher and, thus, the faster CV decoloration observed in the absence of oxygen suggests that a triplet state could be involved in the CV degradation process.

It is worth to be mentioned that CV has been reported to form complexes with humic substances [25]. Formation of similar CV-SBO complexes deserve to be investigated too. Mechanisms involving photosensitizer-pollutant complexes have already been proposed for organic photocatalysts, such as for the photochemical elimination of pesticides in the presence of triphenylpyrylim [26]. In order to gain further insight into this possible mechanistic route, the absorption spectra of CV were recorded in the range 400-800 nm with different amounts of SBO in aqueous solutions at pH = 7. Figure 4 shows that addition of SBO resulted in significant changes in the visible spectrum of CV. Neat CV exhibited an absorption band with a maximum at 583 nm and a shoulder at ca. 540 nm. The presence of increasing SBO concentrations resulted in a decrease of the intensity of the band at 583 nm and affected the maximum/shoulder absorbance ratio. Similar changes in the absorption spectra of CV in the presence of anionic surfactants were attributed to the formation of ion pairs and subsequent dye dimerization [27-29]. The latter phenomenon could be justified by assuming that association of anionic surfactants with the cationic dye results in a decrease of the electrostatic repulsions between the  $=N^+(CH_3)_2$  groups of the dye, thus favoring dimerization of the dye molecules [30]. Also with SBO several species could in principle be formed, and such an occurrence would explain the changes observed in the CV spectrum upon increasing SBO concentration. To better understand the system, the multivariate curve resolution-alternative least square (MCR-ALS) method was chosen as a powerful tool for resolving spectra and concentration profiles of the mixture. This method allows the estimation of changes in concentration and spectra without any previous assumptions. The SVD analysis of the absorption matrix shows that all traces can be reproduced by considering at least four independent components

whose concentration profiles and absorption spectra are shown in Figures 5A and 5B, respectively.

Figure 5A shows that for SBO concentration below  $15 \text{ mg L}^{-1}$ , addition of this material resulted in a decrease of free CV and in the formation of a new species, probably a complex between CV and SBO  $(\text{CV-SBO})_{\text{I}}$ , which shows a blue shift in the UV-vis spectrum. The amount of this intermediate component decreased as the SBO CVT230 concentration increased above  $15 \text{ mg L}^{-1}$ . Another species,  $(\text{CV-SBO})_{\text{II}}$ , was formed and it became predominant above  $40 \text{ mg L}^{-1}$  of SBO. As the concentration of SBO increases, molecular conformation changes and formation of pre-micellar aggregates are expected [31]. Hence,  $(\text{CV-SBO})_{\text{II}}$  might result from the electrostatic interaction between CV and the SBO molecules in the changed solution conformation at this concentration. Figure 5A also shows that at SBO concentrations higher than  $30 \text{ mg L}^{-1}$ , the amount of free CV starts to grow again. The same trend has also been reported for CV in the presence of sodium dodecylsulfate (an anionic surfactant) [32] and it has been attributed by the authors to the solubilization of CV in the surfactant aggregates.

Interestingly, the maximum value of  $k_{\text{app}}$  is reached at SBO contents for which the MCR-ALS model yields the maximum concentration of the species associated with the complex  $(\text{CV-SBO})_{\text{I}}$ . Furthermore, the observed enhancement of the degradation rate of CV in the absence of molecular oxygen (see Figure 3) might imply the involvement of a triplet state of the complex. Nonetheless, SBO are too complex to rule out other mechanistic routes.

### **3.2 Neutral photo-Fenton in the presence of SBO**

The SBO used in this work (CVT-230), contain a significant mineral portion, in particular iron at 0.77 % (w/w) level. This might be of interest as the possibility of

driving a (photo)-Fenton process at circumneutral conditions in the presence of organics able to complex iron, as recently reported for humic acids [15, 16, 32, 33]. To check the possible use of SBO as Fenton additives, solutions containing CV were irradiated under the following experimental conditions: a) in the presence of SBO ( $100 \text{ mg L}^{-1}$ ), b) with  $\text{H}_2\text{O}_2$  ( $8.3 \text{ mg L}^{-1}$ ), c) in the presence of both SBO and  $\text{H}_2\text{O}_2$ . The decoloration rate was fitted to a pseudo-first order kinetics and the apparent rate constant ( $k_{\text{app}}$ ) was determined in all cases (see Table 1). A comparison among the  $k_{\text{app}}$  values shows that although some decoloration was observed with SBO or  $\text{H}_2\text{O}_2$ , the fastest reaction was reached when both species were combined. Experiments with extra amounts of iron(II) ( $5 \text{ mg/L}$ ) were also performed. Table 1 shows that the combination of iron salts and hydrogen peroxide without SBO has scarce effect on the CV decoloration. Similar results are obtained in experiments with  $\text{H}_2\text{O}_2$  in the absence of iron. On the other hand, when SBO was present in the reaction media, the reaction rate was greatly enhanced. A possible explanation of this behavior is that when iron is complexed by the SBO, it is able to participate in a photo-Fenton like process, while in the absence of this substance iron precipitates to form non-active oxides or hydroxides that inhibit the reaction.

Several recent studies support the existence of a modified mechanism for the (photo)-Fenton reaction at circumneutral pH values [32, 34, 35]. Although the nature of the oxidants is still a controversial issue, strong evidence suggests that an oxidant other than  $\text{HO}^\bullet$ , such as the ferryl ion ( $\text{FeO}^{2+}$ ), is formed when the pH is close to neutral values [36]. Moreover, the presence of organic ligands (e.g., EDTA) plays an important role in determining the nature of the oxidant species formed [37]. Since in the present work the experiments have been conducted at neutral pH and probably in the presence of complexed iron, the conditions seem appropriate to gain further insight into the nature of the reactive species. For this purpose, the photodegradation of CV in solutions

containing with SBO, H<sub>2</sub>O<sub>2</sub> and extra amounts of iron was investigated in the presence of: i) 2-propanol, a selective scavenger of hydroxyl radical, which exhibits low reactivity against ferryl ion [34], and ii) phenol, which has been reported to react with both HO<sup>•</sup> and ferryl ion [38]. Table 1 shows that while 2-propanol had a scarce effect on the decoloration rate of CV, in the case of phenol a noticeable inhibition in the reaction was observed. This is compatible with ferryl ion playing a more important role in the process than HO<sup>•</sup>. This hypothesis is in agreement with the above cited observation for neutral photo-Fenton.

#### **4. Conclusions**

Soluble organic biosurfactants (SBO) have been demonstrated to be materials of interest for photochemical wastewater treatment processes, as photosensitizers and/or for the implementation of circumneutral photo-Fenton processes. Hence, further research is needed, in particular: a) to evaluate the potential applicability of SBO for the photoremediation of major occurring environment pollutants, such as pesticides and pharmaceuticals, b) to gain further insight into the reaction mechanism and the species involved in the process, and c) to optimize new promising photoremediation processes.

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## Figure Captions

**Figure 1.** Absorption spectra of aqueous solutions of CV in the presence of SBO obtained at different irradiation times.  $[CV] = 10 \text{ mg L}^{-1}$ ;  $[SBO] = 25 \text{ mg L}^{-1}$ ;  $\text{pH} = 7$ . Inset: Plot of  $\log(A_{\text{max}}/A_{0,\text{max}})$  vs. time obtained for photolysis of CV ( $10 \text{ mg L}^{-1}$ ) in the absence and in the presence of SBO ( $25 \text{ mg L}^{-1}$ ).

**Figure 2.** Apparent bleaching constant ( $k_{\text{app}}$ ) as a function of SBO content.  $[CV] = 10 \text{ mg L}^{-1}$ ;  $\text{pH} = 7$ . The error bars represent standard deviations.

**Figure 3.** Photodegradation of CV ( $10 \text{ mg L}^{-1}$ ) with SBO ( $25 \text{ mg L}^{-1}$ ) in the presence of 2-propanol (0.03 M) and furfuryl alcohol (FFA) (0.001 M). For comparison purpose an experiment without oxygen (purged with  $\text{N}_2$ ) was added.

**Figure 4.** UV-vis spectra of CV with different amounts of SBO.  $\text{pH} = 7$ .

**Figure 5.** A) The concentration profiles at different SBO concentrations and B) the pure spectra of the four components calculated using MCR-ALS.

## Tables

**Table 1.** Apparent pseudo-first order rate constant for the photodecoloration of CV solutions under different conditions.

N° exp.	Conditions	$k_{app} (\times 10^3)$ ( $\text{min}^{-1}$ )
1	CV	$0.58 \pm 0.06$
2	CV + H <sub>2</sub> O <sub>2</sub>	$1.57 \pm 0.06$
3	CV + H <sub>2</sub> O <sub>2</sub> + Fe(II)	$1.49 \pm 0.09$
4	CV + SBO	$1.79 \pm 0.07$
5	CV + SBO + H <sub>2</sub> O <sub>2</sub>	$2.79 \pm 0.11$
6	CV + SBO + H <sub>2</sub> O <sub>2</sub> + Fe(II)	$5.54 \pm 0.30$
7	CV + SBO + H <sub>2</sub> O <sub>2</sub> + Fe(II) + 2-propanol	$5.40 \pm 0.20$
8	CV + SBO + H <sub>2</sub> O <sub>2</sub> + Fe(II) + phenol	$4.09 \pm 0.30$

[CV] = 10 mg L<sup>-1</sup>; [H<sub>2</sub>O<sub>2</sub>] = 8.3 mg L<sup>-1</sup>; [Fe(II)] = 5 mg L<sup>-1</sup>; [SBO] = 100 mg L<sup>-1</sup>; [2-propanol] = 1.8 g L<sup>-1</sup>; [phenol] = 2.8 g L<sup>-1</sup>. pH = 7.0. The error bars represent standard deviations.

Figure  
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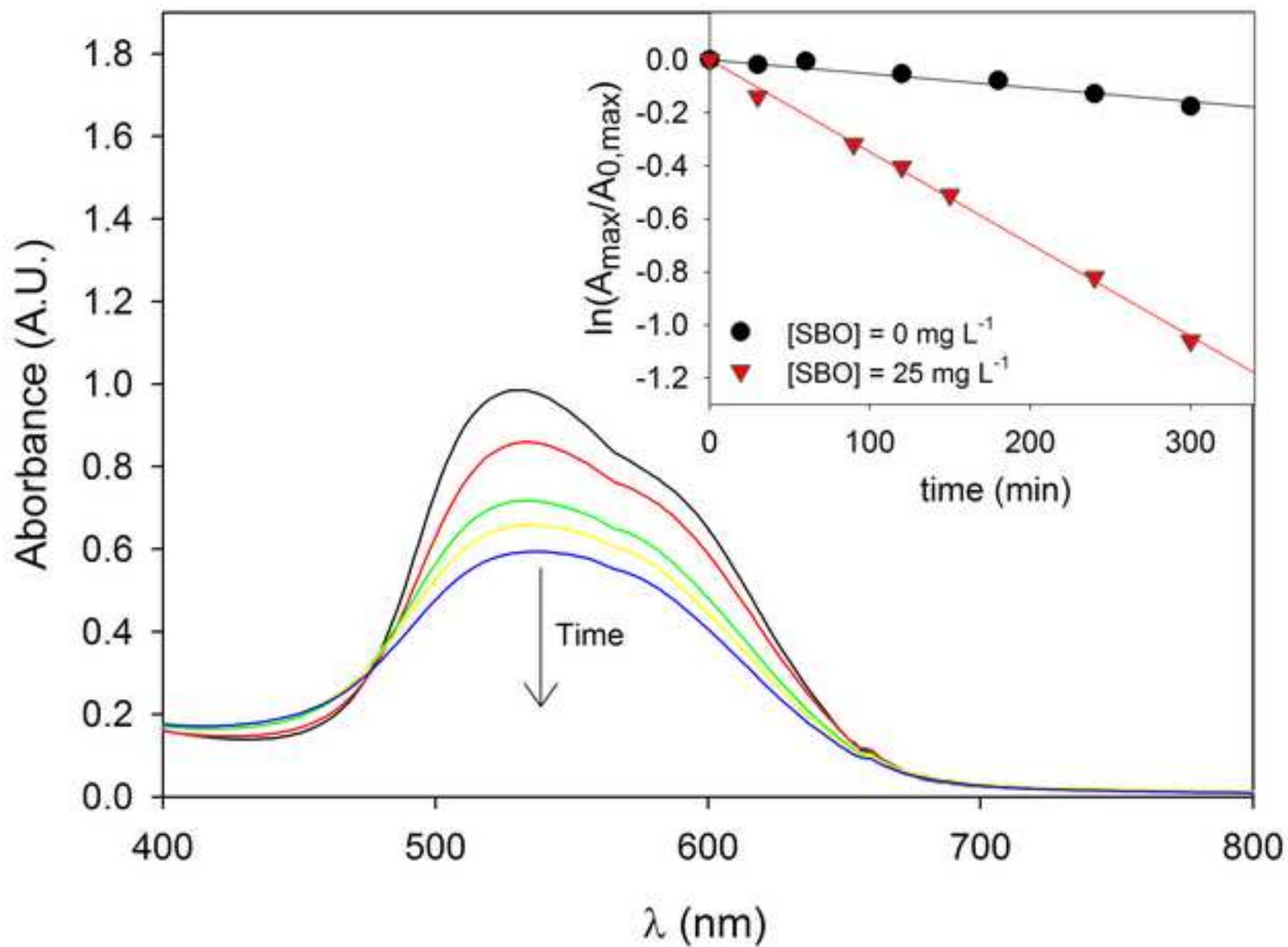


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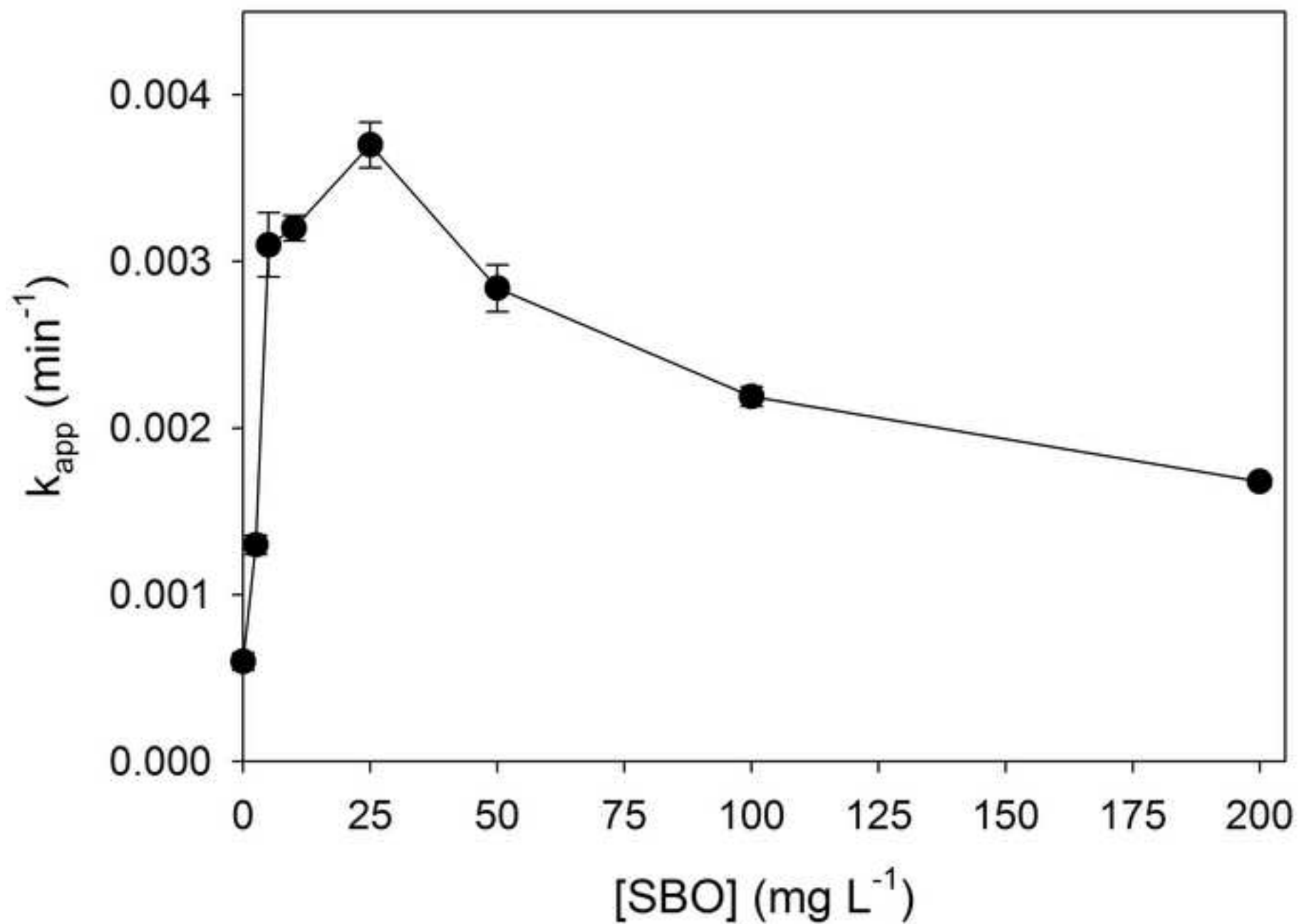


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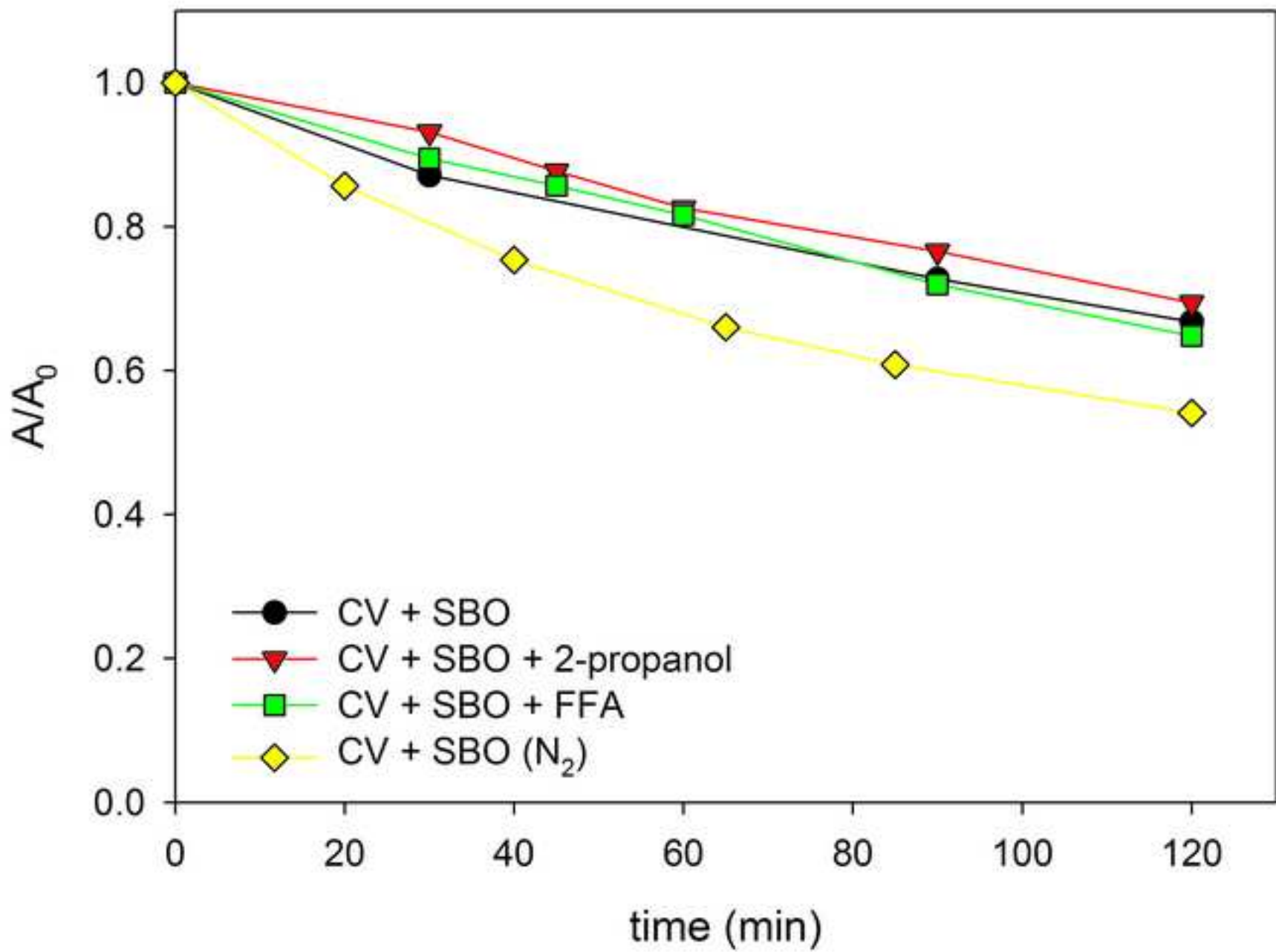


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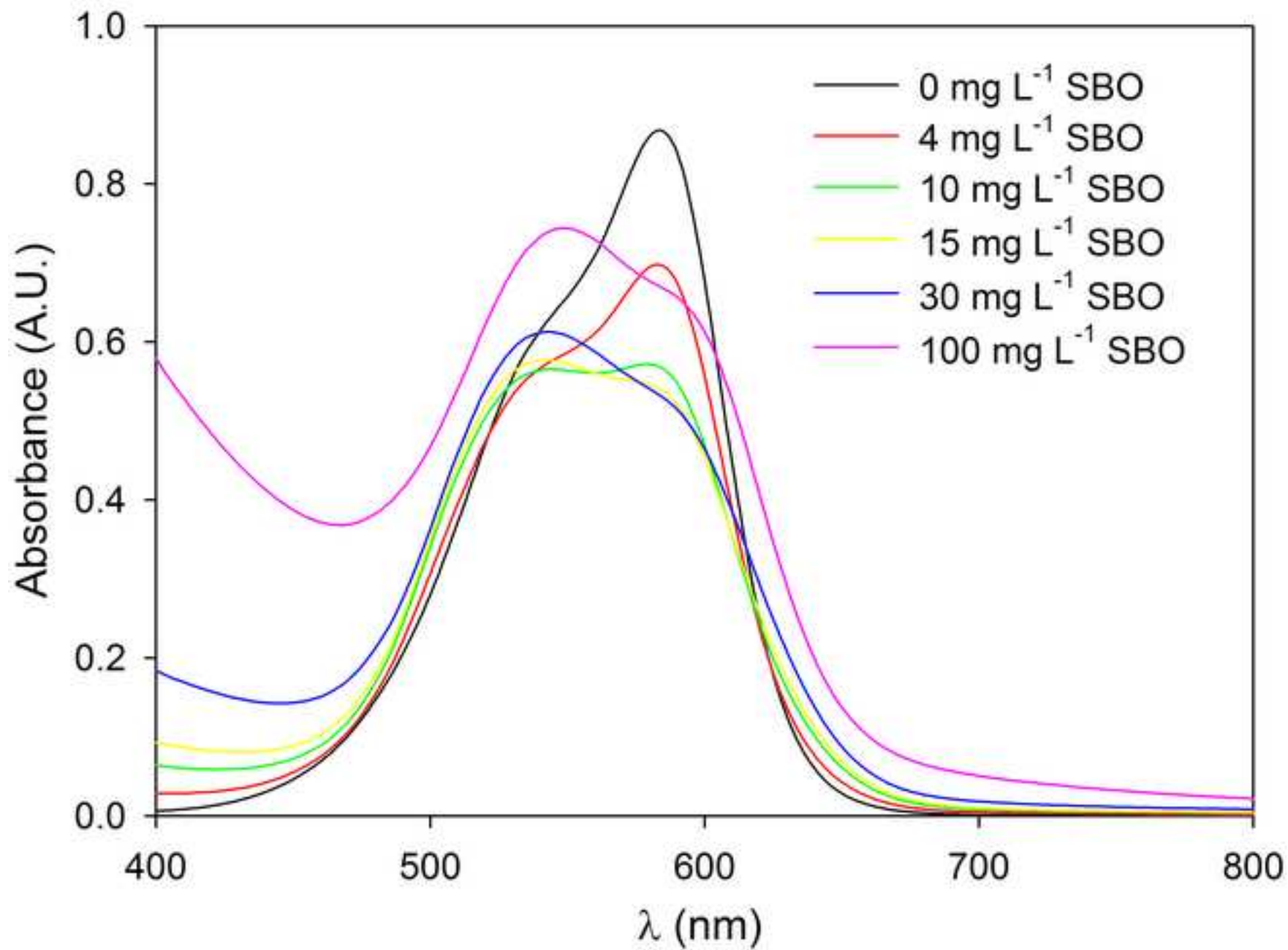


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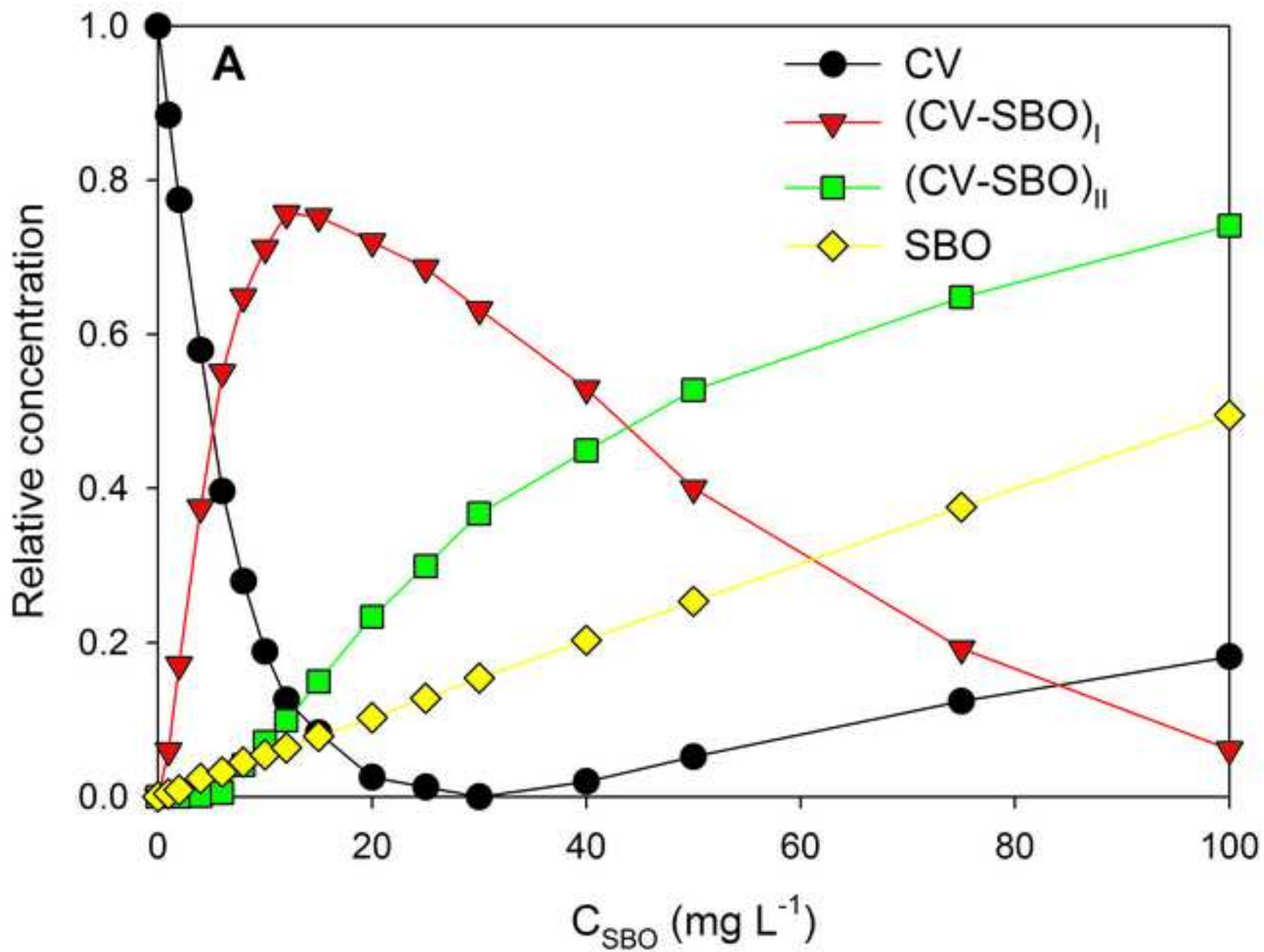




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