

Document downloaded from:

<http://hdl.handle.net/10251/63406>

This paper must be cited as:

Ferre Vilaplana, A.; Perales Rondón, JV.; Feliu, JM.; Herrero, E. (2015). Understanding the effect of the adatoms in the formic acid oxidation mechanism on Pt(111) electrodes. ACS Catalysis. 5(2):645-654. doi:10.1021/cs501729j.



The final publication is available at

<http://dx.doi.org/10.1021/cs501729j>

Copyright American Chemical Society

Additional Information

This document is the Accepted Manuscript version of a Published Work that appeared in final form in ACS Catalysis, copyright © American Chemical Society after peer review and technical editing by the publisher. To access the final edited and published work see <http://dx.doi.org/10.1021/cs501729j>

Understanding the effect of the adatoms in the formic acid oxidation mechanism on Pt(111) electrodes

*Adolfo Ferre-Vilaplana^{‡‡}, Juan Víctor Perales-Rondón[†], Juan M. Feliu[†] and Enrique
Herrero^{†*}.*

[‡]Instituto Tecnológico de Informática, Ciudad Politécnica de la Innovación, Camino de Vera
s/n, E-46022 Valencia, Spain,

^{‡‡}Departamento de Sistemas Informáticos y Computación, Escuela Politécnica Superior de
Alcoy, Universidad Politécnica de Valencia, Plaza Ferrándiz y Carbonell s/n, E-03801 Alcoy,
Spain.

[†]Instituto de Electroquímica, Universidad de Alicante, Apdo. 99, E-03080 Alicante, Spain.

*Corresponding author e-mail: herrero@ua.es

Keywords: formic acid oxidation, platinum, adatom modification, activation energy, DFT.

The engineered search for new catalysts requires a deep knowledge about reaction mechanisms. Here, supported by a combination of computational and experimental results, the oxidation mechanism of formic acid on Pt(111) electrodes modified by adatoms of the p-block is elucidated for the first time. DFT calculations reveal that some adatoms, such as Bi or Pb, have positive partial charge when adsorbed on the bare surface whereas others, Se or S, remain virtually neutral. When the partial charge is correlated with previously reported experimental results for the formic acid oxidation reaction, it is found that the partial positive charge is directly related to the increase in catalytic activity of the modified surface. Further, it is obtained that such a positive partial charge is directly proportional to the electronegativity difference between the adatom and Pt. Thus, the electronegativity difference can be used as an effective descriptor for the expected electrocatalytic activity. Furthermore, this partial positive

charge on the adatom drives the formic acid oxidation reaction, since it favors the formation and adsorption of formate on the adatom. Once adsorbed, the neighboring platinum atoms assist in the C-H bond cleavage. Finally, it is found that most of the steps involved in the proposed oxidation mechanism are barrierless, which implies a significant diminution of the activation barriers when compared to the unmodified Pt(111) electrode. This diminution in the activation barrier is experimentally corroborated for the Bi-Pt(111) electrode, supporting the proposed mechanism.

1. Introduction.

The commercial exploitation of the fuel cell concept depends on the availability of new electrocatalysts capable of providing higher current densities at lower overpotentials. In the past, the search for new and better catalysts has taken place mainly by means of trial and error. However, a much more engineered searching approach, based on knowledge about reaction mechanisms and parameters determining reactivity, is increasingly feasible. A clear example, of the novel approach is the recent development of new catalysts for the hydrogen oxidation reaction. For this rather simple reaction, in which two electrons are exchanged and only a single H-H bond is cleaved, it was experimentally found that the activity is related to the adsorption energy of atomic hydrogen on the catalytic metal surface.¹⁻³ For this reaction, when activity is plotted versus hydrogen adsorption energy typical volcano curves are obtained,⁴ from which the ideal hydrogen adsorption energy giving rise to the maximum activity can be determined. With the advent of powerful computers and methods, hydrogen adsorption energies for a large number of different alloys have been calculated, and new and more effective electrocatalysts whose energy is close to the ideal value have been identified.⁵

A similar strategy has been followed for the oxygen reduction reaction, where the reactivity has been associated to the adsorption energy of different reaction intermediates.⁶

In general, the application of the exemplified engineered approach to the complete oxidation of small organic molecules to CO₂ can become much more complicated. Since several heteroatomic bonds have to be cleaved and, in many cases, even new bonds could be formed, the search for correlations between surface properties and reactivity become much more complex. From all the small organic molecules being investigated as possible fuel, formic acid is clearly the simplest one. The complete oxidation of formic acid to yield CO₂ involves the cleavage of the O-H bond of the carboxylic group and the C-H bond cleavage, giving rise to the exchange of two electrons. The first bond cleavage is rather simple, since it is related to the acid/base equilibrium of the formic acid molecule, whereas the second one requires an active metal surface. Additionally, the oxidation of the formic acid molecule can proceed through a second route in which adsorbed CO is formed through the cleavage of the C-H and the C-OH bonds.^{7,8} Since CO is strongly adsorbed on the electrode surface, its final oxidation to CO₂ it is only achieved at large overpotentials. Thus, CO can be considered as a poisoning intermediate, for this reason the formic acid oxidation route through CO should be avoided. The route yielding CO₂ without forming CO is known as the direct route or the active intermediate route because its reaction intermediates are easier oxidized than adsorbed CO.

From all pure metals, platinum and palladium deploy the highest oxidation activity.⁷ Since the formic acid oxidation reaction involves the interaction of the reactant species with the catalytic surface, a dependence of the reactivity on the surface structure has been found using platinum single crystal electrodes.^{9,10} In general, both described routes are structure sensitive following the same reactivity trend, i.e., the platinum surfaces that are more active for the oxidation through the active intermediate route are also more active for the CO route.^{11,12} To

increase the platinum activity towards the desired route and avoid CO formation, several strategies have been attempted, which involve the modification of platinum surfaces with submonolayers of different adatoms,¹³⁻²⁷ metallic layers,²⁸⁻³⁰ and the preparation of intermetallic compounds with platinum.³¹⁻³⁵ In spite of the fact that, large activity enhancements have been measured for some of the tested electrodes, especially when Bi or Pb is added, there is still no detailed explanation for the observed reactivity increase. Here, , we provide, for the first time, a complete and comprehensive picture of the reaction mechanism of formic acid on Pt(111) surfaces modified with adatoms of the p-block, starting from a critical review of previously reported experimental results, and using a combination of new experimental and computational data. Moreover, a descriptor that correlates previously reported reactivity and physical properties of the adatom has been identified. This descriptor can be used to feed future engineered searching for new catalysts processes.

2. Experimental and computational methods.

Experimental methods.

Platinum single crystal electrodes were oriented, cut and polished from small single crystal beads (2.5 mm diameter) following the procedure described by Clavilier and co-workers.^{36,37} The electrodes were cleaned by flame annealing, cooled down in H₂/Ar and protected with water in equilibrium with this gas mixture to prevent contamination before immersion in the electrochemical cell, as described in detail elsewhere.^{37,38} The voltammetric profiles, and therefore the surface structure of the electrodes, are stable upon cycling provided that oxide formation is avoided. It is known that oxidation/reduction cycles create defects on the electrode surface.^{39,40}

Bismuth adlayers were prepared by the irreversible adsorption technique. The prepared electrode surface is put in contact in a separate vessel with a solution of Bi₂O₃ (10⁻⁵–10⁻⁴ M)

in 0.1 M perchloric acid for a short period of time (5–60 s).⁴¹ Depending on the Bi solution concentration and immersion time, different Bi coverages on the Pt(111) electrode can be obtained. The modified electrode was then rinsed with water and immersed again in the electrochemical cell. The coverage of irreversibly adsorbed bismuth was calculated using the charge of the new peak appearing in the voltammogram in 0.5 M H₂SO₄ at 0.67 V. Bismuth coverage is given as the number of Bi atoms per Pt surface atoms and has been calculated according to the procedure described in reference.²⁶ Experiments were carried out in a classical two-compartment electrochemical cell deaerated by using Ar (N50, Air Liquide in all gases used), including a large platinum counter electrode and a reversible hydrogen (N50) electrode (RHE) as reference. For the determination of the activation energy, the electrochemical cell was immersed in a water bath to control the temperature in a range between 278 and 333 K. The reference electrode was kept at room temperature (298 K) and all the measured potentials are referred to a RHE electrode at 298 K. For that reason, the measured potentials were corrected with the thermodiffusion potential using the procedure explained in reference⁴². All the potentials are quoted vs. the RHE at 298 K unless otherwise stated. Solutions were prepared from sulfuric acid, perchloric acid, sodium perchlorate, acetic acid, formic acid (Merck suprapur in all cases) and ultrapure water from Elga. The cleanliness of the solutions was tested by the stability of the characteristic voltammetric features of well-defined single crystal electrodes.

The potential program for the transients was generated with an arbitrary function generator (Rigol, DG3061A) together with a potentiostat (eDAQ EA161) and a digital recorder (eDAQ, ED401). To avoid any interference of the diffusion of formic acid in the reaction rate, stationary conditions were attained by using a hanging meniscus rotating disk configuration at 900 rpm (controlled by a Radiometer CTV 101).

Computational methods.

All DFT calculations were carried out using numerical basis sets,⁴³ semicore pseudopotentials⁴⁴ (which include scalar relativistic effects) and the RPBE functional⁴⁵ as implemented in the Dmol³ code.⁴⁶ Implicit solvation effects were taken into account by the COSMO model.⁴⁷ The effects of non-zero dipole moments, in the supercells, were cancelled by means of external fields.⁴⁸

The Pt(111) surface was modeled by means of a periodic supercell comprising 36 Pt atoms (four layers of metal atoms) and a vacuum slab of 20 Å. The bottom 18 Pt atoms were frozen in their bulk crystal locations. The remaining 18 Pt atoms were completely relaxed joint to the adsorbates. The shortest distance between periodic images was 8.51 Å.

Optimal configurations (adsorbent/adsorbate, reactants, products and transition states) were searched for using numerical basis sets of double-numerical quality. During this phase of the calculations, Brillouin zones were sampled, under the Monkhorst-Pack method,⁴⁹ using grids corresponding to distances in the reciprocal space of the order of 0.05 1/Å, and convergence was facilitated introducing 0.005 Ha of thermal smearing.

Assuming the previously optimized configurations, binding energies, reaction energies and barriers were estimated using numerical basis sets of double-numerical quality plus polarization. In this case, Brillouin zones were sampled, also under the Monkhorst-Pack method, but using grids corresponding to distances in the order of 0.04 1/Å, and convergence was facilitated introducing 0.005 Ha of thermal smearing. Moreover, energies were extrapolated to 0°K. Finally, molecular dynamic simulations were run, under the NVE thermodynamic ensemble, using the same numerical setup used for optimization calculations and 1 femtosecond of time step.

3. Results.

Critical review of previous experimental results.

Searching for increasing the activity of the bare surface towards the formic acid oxidation reaction, the electrocatalytic behavior of Pt(111) electrodes modified with different adatoms has been extensively investigated. Among them, the clearest example giving rise to a significant improvement in the electrocatalytic activity is probably that of bismuth.^{15,26} For this modified electrode, as the adatom coverage increases, two important effects can be observed: the complete inhibition of the route through CO and the significant increase of the oxidation currents measured at low potentials. In fact, it has been reported that the dehydration path yielding CO is completely inhibited even for very low coverages.¹⁸ Given that on the Pt(111) bare surface CO formation takes place exclusively on defect sites^{50,51} and that the initial deposition stages of the bismuth adlayer also occur on the defects, very small amounts of adatom are able to completely block the CO formation reaction.⁵² However, it should be stressed that this inhibition is not linked to a change in the oxidation mechanism, since no major catalytic effects are measured at small coverages.¹⁵ Once all the defect sites have been covered, deposition starts on the well-ordered (111) domains. From these coverage levels forward, a significant increase in the catalytic activity towards CO₂ can be observed until the coverage value of 0.26 is reached.^{15,26} At this optimal coverage value, measured currents for the direct oxidation of formic acid on the modified electrode at 0.5 V are ca. 15 times higher than those measured for the bare surface.^{12,26} For higher coverages, the activity diminishes until the maximum stable coverage of 0.33 is obtained. At this latter coverage value, the electrocatalytic activity is almost negligible,¹⁵ which may be explained by the fact that, at such a coverage level, all platinum sites have been covered. From the above presented results, two conclusions arise: i) the presence of Bi on the (111) terraces catalyze the direct oxidation of formic acid to CO₂ and ii) the availability of Pt sites seems to be a requirement

for the oxidation process to take place. In fact, it has been proposed that a Pt site close to a Bi adatom would be the most catalytic active site for the considered oxidation reaction.⁵³ Similar conclusions have been reached for stepped surfaces with (111) terraces.^{24,25} Additionally, it has been shown by STM that Bi is distributed randomly over the surface, since the sizes of Bi islands on the surface are below 1 nm for low and medium coverages.⁵⁴ Assuming such a random distribution of the Bi adatom on the Pt(111) surface, a statistical model for the adatom modified surface suggests that the Bi-Pt ensemble is the most active site for the considered reaction.⁵³

For the formic acid oxidation reaction, other adatoms, such as Pb, Te, As or Sb, on Pt(111) electrodes, behave as bismuth, although with different catalytic activities.^{17,19,21-23} Also intermetallic Pt-Pb nanoparticles display a very high electrocatalytic activity towards the considered reaction, in a similar way to that observed for Pt-Bi intermetallic nanoparticles.³³⁻³⁵ Actually, the Pb-Pt(111) system displays the highest catalytic activity for the target reaction, though Pb as adatom on the Pt(111) electrode has some stability problem, since it readily dissolves at potentials higher than 0.5 V.¹⁹ On the other hand, certain adatoms, such as Se or S, deposited on the Pt(111) electrode, do not give rise to any catalytic enhancement for the studied reaction.^{20,53}

When the catalytic activities towards the target reaction of all the investigated adatom modified Pt(111) electrodes are compared at 0.5 V, the following order can be established: Pb-Pt(111) > Bi-Pt(111) > Sb-Pt(111) > Te-Pt(111) > As-Pt(111) > Pt(111) > Se-Pt(111) = S-Pt(111). It should be emphasized that, from a practical point of view, the highest power output is always sought out for a fuel cell. So, when the relative performance of different electrodes is evaluated, comparison of currents at constant potential is the relevant figure. Current at peak potential is not a relevant figure for comparison, because, in most of the cases, peak potential depends on the deactivation mechanism of the surface, which can be

affected by the presence of very different adsorbates such as adatoms or anions from the electrolyte.

Adatoms effect on the Pt(111) surface.

In order to understand the effect of adatoms on the bare surface, electronic structure calculations, modeling different adatom modified Pt(111) surfaces, were carried out using DFT. Adsorption energies of adatoms on the Pt(111) surface, both for fcc and hcp sites, are summarized in table 1. The adsorption energy of a Pt atom is also included for comparison in order to distinguish the effects related to the different chemical nature of the adatom from those related to the geometrical position of the adatom. Adatom adsorption processes give rise to electron density redistributions which, are displayed in figure 1 for fcc configurations, being the results corresponding to hcp sites very similar. As can be inferred from figure 1, for some adatoms, such as Pb or Bi, a partial positive charge is located on the adatom, whereas the compensating negative charge is smeared out over the neighboring Pt sites. However, for others, mainly Se and S, the electron density redistribution is very small and both the adatom and the surface remain virtually neutral. To quantify the electron density changes observed upon the adsorption of the adatom, partial charges on the adatoms were estimated both under the Mulliken and Hirshfeld formalisms.^{55,56} For the fcc configuration, such partial charges are summarized in table 2. Similar values have been previously found for the Sb-Pt(111) system,⁵⁷ and similar trends have been reported for the adsorption of Bi and S on small Pt₁₀ clusters.⁵⁸ From the comparison of the partial charges series with the catalytic activity sequence listed at the end of the critical review section, it can be concluded that there exist a clear correlation between the partial charge of the adatom and the catalytic activity of the system. Higher partial positive charge on the adatom leads to higher electrocatalytic activity towards the formic acid oxidation. Moreover, adatoms that do not have any positive partial

charge such as Se or S do not exhibit any catalytic enhancement. Thus, it seems clear that the measured catalytic activity increase is related to the partial positive charge appearing on the adatom.

Finally, for the investigated adatoms, Pauling electronegativities are also included in table 2. From this table, a correlation between electronegativities and partial charges can be established. In fact, a linear relationship between the partial charge on the adatom (δ) and the electronegativity difference between the adatom and platinum (ΔE_N) is obtained in figure 2. Additionally, it can also be observed that, for zero electronegativity difference, both models intercept the same 0.05 e value. Such a circumstance is consequence of the Smoluchowski effect.⁵⁹ This effect predicts some positive charge on the upper part of the defects of a surface, due to a redistribution of the charge that is not able to follow the geometrical profile. In fact, as can be observed in table 2, a Pt adatom would have a positive partial charge between 0.05 and 0.08, depending of the formalism, consequence of the aforementioned effect. Thus, the charge appearing on the adatom can be explained by the electronegativity difference between the adatom and Pt. Given the correlation between partial charge and catalytic activity, it can be concluded that, the electronegativity difference between the adatom and Pt can be used as a descriptor for the expected catalytic enhancement towards the formic acid oxidation reaction on the adatom modified Pt(111) electrode.

Formic acid oxidation mechanism.

Being established that there exists a clear correlation between the positive charge located on the adatom upon adsorption on the Pt(111) surface and the formic acid oxidation currents measured experimentally, the precise effect of the adatom on the formic acid oxidation process was investigated. Several possibilities were considered. Formic acid (or formate) could interact first with the positively charged adatom and then reacting to yield CO₂.

Alternatively, formic acid (or formate) could adsorb on the Pt atoms close to the adatom and from that state evolves yielding the final product. Using experimental and DFT results, here it is demonstrated that the most favorable mechanism proceeds through adsorption on the adatom, that is, the adatom plays an active role in the oxidation mechanism.

For mechanism in which the adatom plays an active role, the effect of the adatom partial positive charge on the formic acid molecules approaching the surface was considered first. Physisorption energies of a hydrated formic acid molecule adsorbed near the different adatoms, in two configuration (with the carbonyl group and the hydroxyl group close to the adatom), are summarized in table 3. In spite of that, under classical functionals, DFT is not the more accurate method in order to capture relatively weak interactions, the results presented in table 3 suggest that the physisorption of hydrated formic acid on the adatom modified surfaces is significant, and that the interactions are strong enough as to be sufficiently captured by a classical functional. Additionally, it can be observed that the estimated physisorption energies are proportional to the positive charge located on the adatoms given in table 2. Also, the data in table 3 shows that, for the adatoms with a large partial positive charge (i.e., Bi and Pb), the interaction with the carbonyl group is more favorable, as was previously shown for Sb.⁵⁷ For the rest of adatoms, the differences in energy between both configurations are within the error of the calculations. As expected, the distances between the formic acid molecule and the adatom become shorter as the positive charge in the adatom increases (figure 3), since there is a clear interaction of the partial negative charge of the carbonyl group and the positive charge of the adatom. Moreover, the physisorption of the formic acid molecule on the adatom also involves some structural changes in the hydrated formic acid ensemble, with changes in the distances of O-H bond and the hydrogen bond with respect to those calculated in absence of the surface.

From the described physisorbed configuration, the formic acid molecule may evolve to chemisorbed formate by deprotonation towards the solution. The positive charge on the adatom (figure 1) would be the driving force of such a formate chemisorption process in two different ways. On one side, given that positive charges interact favorably with anions, the adatom would become as an attractor for the conjugated base of the formic acid. On the other hand, the positive charge on the adatom would drive the chemisorption process of formate towards the adatom instead of platinum. In this solution pH, that is, when $\text{pH} < \text{pK}_a$, the average number of acid molecules at any given time is larger than those of the conjugated base, but there is a continuous and dynamic exchange between both forms, as in any other equilibria. Each time a formic acid molecule close to the adatom initiates the deprotonation process, the positive charge on the adatom would increase the probability that the molecule can be finally deprotonated and adsorbed on the adatom. Actually, from the formic acid physisorbed configuration, deprotonation and formate chemisorption processes on the adatom would take place simultaneously, as DFT calculations corroborate. Due to the large positive charge located on some adatoms (see table 2), adsorption of formate on the adatoms is energetically favorable and it occurs in a barrierless process, where formate would be bonded to adatoms through an oxygen atom (figure 4A). Moreover, running quantum mechanics molecular dynamic simulations, it was found that, in this chemisorbed mode, the formate fragment can rotate freely using the adatom-oxygen and the contiguous oxygen-carbon bonds as rotation axis. Two major configurations were identified, in which there is a clear interaction with the surface: with the C-H bond pointing to the surface (C-H down configuration, figure 4A) and that with the C-H bond in the opposite direction (C-H up configuration). For all the adatoms, the energy difference between both configurations is below 0.1 eV, which clearly indicates that the molecule can rotate freely, because all other configurations have energies comprised between these two extreme values.

Finally, in order to yield CO₂, from the assumed formate chemisorption state on the adatom, the C-H bond should be cleaved. For this last step of the whole process, the identified C-H down configuration is the relevant one. Taking into account the large affinity of platinum for hydrogen, it is possible that C-H bond cleavage takes place from such a configuration, in which the hydrogen atom is close to a Pt site. Moreover, it has been shown that the activation energy for the cleavage of the C-H bond on a Pt(111) surface from a C-H down configuration is significantly smaller than that measured for a C-H up configuration.⁶⁰ Thus, the energetics of the C-H bond cleavage schematized in figure 4 was computed for all the adatoms. For all cases, the resulting energy of the process was found to be favorable, and in most cases barrierless (table 4), which indicates that the evolution of the adsorbed fragment to yield CO₂ is a favorable process for the adatoms that have a positive partial charge when adsorbed on Pt(111) surfaces. The total energy of this last step is less favorable as the partial positive charge of the adatom increases, although, in all cases, the process is clearly favorable. This is just a consequence of the difference in the adsorption energy of formate on the adatom. As the global process should have the same energy independently of the nature of the adatom, the adsorption energy of formate is more favorable as the partial positive charge on the adatom is larger, so that the energy released in the last step diminishes.

Having established that the mechanism under which the adatom adsorb the HCOO fragment and the neighboring Pt atom helps in the cleavage of the C-H bond is almost a barrierless mechanism (the only expected energy barrier is deprotonation of the physisorbed formate molecule), the alternative mechanism involving the adsorption of formate on the Pt sites close to the adatom should be analyzed. In all the reported mechanisms for Pt surfaces, monodentate and bidentate adsorbed formate are key intermediate species.^{61,62} Thus, the adsorption energy of monodentate and bidentate formate (HCOO_m and HCOO_b, respectively) adsorbed near the adatom on the Bi modified Pt(111) surface and on the unmodified Pt

surface were estimated. From all the considered adsorption processes, the adsorption of HCOO_b on the unmodified surface is the most stable. Table 5 reports those energies using the adsorption energy of HCOO_b on the unmodified Pt(111) surface as reference. As can be seen, the adsorption of HCOO_b on the Pt sites close to the Bi is ca. 0.17 eV less favorable than that estimated for the bare Pt(111) surface. This result is in very good agreement with the experimental behavior of adsorbed formate on Bi modified Pt(111) surfaces inferred from fast voltammetry.⁶³ The onset for formate adsorption is displaced ca. 0.2 V for $\theta_{\text{Bi}}=0.12$, which is a coverage very similar to that corresponding to the calculations. Also the difference in adsorption energy between HCOO_m and HCOO_b for the unmodified Pt(111) surface are in good agreement with reported results.⁶¹ In fact, the less favorable adsorption of formate on the adatom modified surfaces can be also considered a consequence of the charge redistribution originated by the adsorption of the adatom. Due to the positive charge on the adatom, the Pt atoms have a partial negative charge, which hinders the adsorption of nucleophilic species, such as formate.

For a mechanism in which the adatom only modifies the energy of the neighboring Pt sites, that is, it is not involved directly in the catalysis, the lower adsorption energy of formate on the Pt sites close to the adatom will lead to higher barriers and therefore to lower current densities, as the results on the Pt(111) surface demonstrates.⁶¹ It should be stressed that the calculated energy barriers for the whole mechanism on the unmodified Pt(111) surface are well above 0.5 eV,⁶¹ whereas the mechanism in which formate is adsorbed on the adatom is almost barrierless.

The participation of the adatom in the mechanism by aiding in the C-H cleavage when the HCOO fragment is adsorbed in a neighboring Pt can also be discarded. It has been shown that the cleavage of the C-H bond in the HCOO_b configuration on Pt(111) has a very high barrier (ca 1 eV).⁶⁰ In this configuration, due to geometrical reasons, the interaction of the C-H bond

with the adatom will be small, so that the activation barrier will continue to be very high. Lower activation barriers for the C-H cleavage on unmodified Pt(111) surfaces have been calculated from the HCOO_m configuration.⁶¹ However, the energy difference between the HCOO_b and the HCOO_m adsorption modes in Bi-Pt(111) surfaces is 0.27 eV, and the energy barrier for the transition will be higher. In comparison with mechanism where the adatom adsorbs the HCOO fragment, this process is less favorable even without considering the barrier for the C-H bond cleavage, since it has already higher barriers. Moreover, the energy barriers for the cleavage of the C-H bond from HCOO_m are not expected to be negligible. In the most favorable mechanism, the partial positive charge of the adatom favors the adsorption of the anion and the affinity of platinum for hydrogen adsorption helps in C-H bond cleavage. In the mechanism initiated by the adsorption on platinum, formate is adsorbed in a less favorable site (due to the negative charge compensating the positive charge located on the adatom) and the cleavage of the C-H bond is occurring on an adatom with lower hydrogen affinity. As a conclusion, the mechanism involving the adsorption of HCOO fragments on Pt cannot explain the enhanced catalytic activity observed.

In summary, the proposed oxidation mechanism of formic acid on adatom modified Pt(111) surfaces in which formic acid interacts first with the adatom is consistent with the experimental observation that adatoms catalyzes the oxidation process, being the adatom-Pt ensemble the most active site. In fact, the above described computational results indicate that the catalysis occurs through a bifunctional mechanism, in which distinct roles are played by the adatom and a Pt site close to the adatom forming an adatom-Pt ensemble. Adatoms having lower electronegativity than platinum give rise to a partial positive charge on the adatom when adsorbed. The positive charge on the adatom attracts the formic acid molecule towards the adatom, and a physisorption phenomenon takes place. From such a physisorbed state, formic acid can deprotonate to give the formate anion, which is immediately chemisorbed on

the adatom. Once chemisorbed, the formate fragment rotates freely around the adatom-O and the contiguous O-C bonds until a sufficiently favorable C-H down configuration is reached. Finally, from such a configuration, the C-H bond is cleaved yielding CO₂ as a product. According to the computed mechanism, all the identified steps are virtually barrierless except the deprotonation of the formic acid molecule, which clearly has some barrier involved at acidic pH values.

Experimental verification of the proposed mechanism.

With the goal in mind of verifying additional aspects of the proposed formic acid oxidation mechanism, new voltammetric experiments on the Bi-Pt(111) system were conducted for three different bismuth coverages and the bare surface (figure 5). In absence of formic acid, three different regions, corresponding to three different processes occurring on the surface, can be identified in the voltammetric profiles for the modified electrodes in sulfuric acid solutions (figure 5A). At 0.06 V, the platinum sites are covered by adsorbed hydrogen, which is desorbed between this potential and 0.2-0.3 V depending on the coverage. As aforementioned, between 0.3 and 0.5 V, sulfate is adsorbed on the available Pt sites and the adsorption/desorption process give rise to the signal in this region. Finally, the peak centered at 0.62 V is related to the OH adsorption on the bismuth adatoms.

The same three different regions can be also identified in the voltammogram recorded for the oxidation of formic acid (figure 5B). In the region where hydrogen is adsorbed, no oxidation currents are detected. The onset for the oxidation of formic acid coincides with the final desorption of hydrogen from the Pt sites. It should be highlighted that the displacement of the onset for the formic acid oxidation towards lower potential values as the adatom coverage increases is parallel to the displacement of the final stages of the hydrogen desorption. This fact is in agreement with the oxidation mechanism proposed by DFT. In the

final step of the process, the cleavage of the C-H bond yields an adsorbed H atom on Pt, which should be desorbed for the reaction to continue. Since the oxidation reaction is not possible in a hydrogen covered surface, the reaction can only proceed at potentials where hydrogen desorption is a spontaneous process. From the onset potential, the oxidation current increases up to 0.5 V, where it reaches a maximum and, at 0.6-0.62 V, the current abruptly diminishes to reach very low values. The abrupt diminution coincides with the potential where the adsorption of OH occurs on the Bi adatom. This fact is also in agreement with the proposed mechanism, since the adsorption of OH will hinder the adsorption of formate, blocking the reaction. Thus, it can be proposed that H adsorption on platinum and OH adsorption on bismuth act as inhibitors of the reaction.

As aforementioned, our DFT results suggest that the activation barriers on Pb or Bi modified Pt(111) electrodes would be clearly lower than those operating on the bare surface. Thus, a significant diminution of the experimental activation energy for the modified electrodes should be measured. Given that, on these modified electrodes, the formic acid oxidation reaction occurs exclusively through the active intermediate route, voltammetry can be used directly to determine the activity through this route. Voltammograms at five temperatures for a Bi-Pt(111) electrode with $\theta_{\text{Bi}}=0.12$ are displayed in figure 6. As can be seen, the hysteresis between the positive and negative scans is very small, which confirms that the formation of CO is negligible. In fact, currents in the positive scan direction are larger than those measured in the negative one. From these measurements, activation energies can be calculated from Arrhenius plots as shown in reference ⁶⁴. Such activation energies are plotted and compared to that measured for the Pt(111) electrode in figure 7. For the unmodified Pt(111) electrode, data has been taken from reference,⁶⁴ in which the activation energy has been calculated using pulsed voltammetry to separate the contribution from the active intermediate route from those related to the CO formation. As shown in figure 7, a

clear diminution of the experimental activation energy is observed as the bismuth coverage increases. For the modified electrodes, three different regions can be identified in the activation energy plot vs. electrode potential. From the onset of the oxidation (ca. 0.2-0.3 V depending on the adatom coverage, see figure 5B), the activation energy diminishes until a minimum value is reached at ca. 0.5 V. Between 0.5 and 0.7 V, the region in which the redox process of the Bi adatom takes place, there is a sudden change in the activation energy. As aforementioned, at ca. 0.6 V, there is an abrupt diminution of the current linked to the adsorption of OH on the adatom that blocks the adsorption of formate. Since this process is affected by temperature,⁶⁵ the exact potential where the sharp diminution occurs is also influenced by temperature. Thus, the apparent activation energy measured in this region is then a consequence of the effect of the temperature in the OH adsorption process on the adatom and is not directly related to the formic acid oxidation reaction. At higher potentials, the currents are very low and the measured activation energy has a significant error.

In the region between 0.3 and 0.5 V, the experimental activation energy contains the contribution of Pt-Bi ensembles where the reaction proceeds through the mechanism proposed in this manuscript and the Pt-Pt ensembles where the reaction proceeds according to the mechanism proposed in ⁶⁴. The experimental activation energy for the Pt(111) electrode with $\theta_{\text{Bi}}=0.28$ is 21 kJ mol⁻¹, whereas the corresponding activation energy on the unmodified Pt(111) electrode at the same potential is 42 kJ mol⁻¹. This fact implies that the activation energy has been reduced to a half and the measured currents have increased more than an order of magnitude.

According to the DFT calculations, even lower activation energy values should have been expected. In a mechanism where activation barriers are small, and therefore, there is no a clear rate determining step, the experimental activation energy should contain the contributions from all the steps. In this mechanism two steps clearly contribute. The first one

is the deprotonation of the formic acid to yield formate, since it is a step which has an unfavorable energy. Assuming that the acid constant of formic acid in the interphase is the same than that measured in the bulk, the corresponding free energy for the deprotonation process in a 0.5 M H₂SO₄ solution (pH close to 0.3) is ca. 9 kJ mol⁻¹. At low potentials, there is an additional contribution to the activation energy from the hydrogen desorption process required to liberate the adsorption site and to reestablish the surface to the initial conditions. The observed diminution of the activation energy between 0.3 V and 0.5 V can be related to this process. Thus, as the potential increases, hydrogen desorption is more favorable, diminishing all the measured activation energy of the reaction. Probably at 0.5 V the contribution of this step is negligible.

The measured activation energy is higher than those values. However, it should be taken into account that the experimental values contains not only information on the activation energy of the whole reaction on the Pt-Bi ensemble. As aforementioned, Pt-Pt ensembles also contributes to the measured activation energy although their number at high bismuth coverages is small.⁶³ Additionally, the dependence with the temperature of all the adsorption processes that take place in the mechanism or are competing for the adsorption sites also contribute to the measured activation energy, as shown for the oxidation on the unmodified Pt(111) electrode.⁶⁴ All these factors increase the measured activation energy.

4. Conclusions.

Experimental results evidence that certain adatoms increase the catalytic activity of Pt(111) electrodes towards the formic acid oxidation reaction. Here, supported by a combination of computational and experimental results, an explanation for such an observation has been provided for the first time. More specifically, the oxidation mechanism of formic acid on Pt(111) electrodes modified by adatoms of the p-block would have been

revealed. The charge redistribution occurring as a consequence of the adatom adsorption process drives the oxidation process. A direct correlation between the positive charge located on the adatom and the activity of the modified electrode has been found. Moreover, such partial positive charge is proportional to the difference in electronegativity between the adatom and platinum, and therefore the electronegativity difference is an excellent descriptor of the expected electrocatalytic activity of the modified surface towards the reaction. Additionally, this partial positive charge is the key factor that drives the oxidation mechanism. It favors the physisorption of formic acid near the adatom and also its deprotonation and further chemisorption, on the adatom, in formate form. In the chemisorbed formate state, the HCOO fragment can rotate freely and, when the H atom is close to a neighboring platinum atom, the cleavage of the C-H bond occurs yielding a CO₂ molecule and an adsorbed hydrogen atom. The final step in the process can be considered to be the desorption of the H atom. Thus, the here proposed oxidation mechanism can be described as a bifunctional mechanism, in which adatoms favor the formation and adsorption of formate and platinum is in charge of the cleavage of the C-H bond. Additionally, for the proposed mechanism, a clear diminution of the activation energy with respect to that operating on the unmodified electrode is computationally predicted and experimentally corroborated.

ACKNOWLEDGMENTS

This work has been financially supported by the MINECO (Spain) (project CTQ2013-44083-P) and Generalitat Valenciana (project PROMETEOII/2014/013).

References.

- (1) Parsons, R. *Trans. Faraday Soc.* **1958**, *54*, 1053-1063.
- (2) Conway, B. E.; Bockris, J. O. *J. Chem. Phys.* **1957**, *26*, 532-541.
- (3) Trasatti, S. *J. Electroanal. Chem.* **1972**, *39*, 163-184.
- (4) Trasatti, S. *J. Chem. Soc. Faraday Trans. I* **1972**, *68*, 229-236.

- (5) Greeley, J.; Jaramillo, T. F.; Bonde, J.; Chorkendorff, I. B.; Norskov, J. K. *Nat. Mater.* **2006**, *5*, 909-913.
- (6) Stephens, I. E. L.; Bondarenko, A. S.; Gronbjerg, U.; Rossmeisl, J.; Chorkendorff, I. *Energy Environ. Sci.* **2012**, *5*, 6744-6762.
- (7) Parsons, R.; Vandernoot, T. *J. Electroanal. Chem.* **1988**, *257*, 9-45.
- (8) Willsau, J.; Heitbaum, J. *Electrochim. Acta* **1986**, *31*, 943-948.
- (9) Clavilier, J.; Parsons, R.; Durand, R.; Lamy, C.; Leger, J. M. *J. Electroanal. Chem.* **1981**, *124*, 321-326.
- (10) Sun, S. G.; Clavilier, J.; Bewick, A. *J. Electroanal. Chem.* **1988**, *240*, 147-159.
- (11) Grozovski, V.; Climent, V.; Herrero, E.; Feliu, J. M. *ChemPhysChem* **2009**, *10*, 1922-1926.
- (12) Grozovski, V.; Climent, V.; Herrero, E.; Feliu, J. M. *Phys. Chem. Chem. Phys.* **2010**, *12*, 8822-8831.
- (13) Watanabe, M.; Horiuchi, M.; Motoo, S. *J. Electroanal. Chem.* **1988**, *250*, 117-125.
- (14) Clavilier, J.; Fernández-Vega, A.; Feliu, J. M.; Aldaz, A. *J. Electroanal. Chem.* **1989**, *261*, 113-125.
- (15) Clavilier, J.; Fernández-Vega, A.; Feliu, J. M.; Aldaz, A. *J. Electroanal. Chem.* **1989**, *258*, 89-100.
- (16) Fernández-Vega, A.; Feliu, J. M.; Aldaz, A.; Clavilier, J. *J. Electroanal. Chem.* **1989**, *258*, 101-113.
- (17) Fernández-Vega, A.; Feliu, J. M.; Aldaz, A.; Clavilier, J. *J. Electroanal. Chem.* **1991**, *305*, 229-240.
- (18) Herrero, E.; Fernández-Vega, A.; Feliu, J. M.; Aldaz, A. *J. Electroanal. Chem.* **1993**, *350*, 73-88.
- (19) Xia, X. H.; Iwasita, T. *J. Electrochem. Soc.* **1993**, *140*, 2559-2565.
- (20) Llorca, M. J.; Herrero, E.; Feliu, J. M.; Aldaz, A. *J. Electroanal. Chem.* **1994**, *373*, 217-225.
- (21) Herrero, E.; Llorca, M. J.; Feliu, J. M.; Aldaz, A. *J. Electroanal. Chem.* **1995**, *394*, 161-167.
- (22) Lei, H. W.; Hattori, H.; Kita, H. *Electrochim. Acta* **1996**, *41*, 1619-1628.
- (23) Climent, V.; Herrero, E.; Feliu, J. M. *Electrochim. Acta* **1998**, *44*, 1403-1414.
- (24) Smith, S. P. E.; Ben-Dor, K. F.; Abruna, H. D. *Langmuir* **1999**, *15*, 7325-7332.
- (25) Smith, S. P. E.; Abruna, H. D. *J. Electroanal. Chem.* **1999**, *467*, 43-49.
- (26) Maciá, M. D.; Herrero, E.; Feliu, J. M. *J. Electroanal. Chem.* **2003**, *554*, 25-34.
- (27) Buso-Rogero, C.; Perales-Rondon, J. V.; Farias, M. J. S.; Vidal-Iglesias, F. J.; Solla-Gullon, J.; Herrero, E.; Feliu, J. M. *Phys. Chem. Chem. Phys.* **2014**, *16*, 13616-13624.
- (28) Llorca, M. J.; Feliu, J. M.; Aldaz, A.; Clavilier, J. *J. Electroanal. Chem.* **1994**, *376*, 151-160.
- (29) Baldauf, M.; Kolb, D. M. *J. Phys. Chem.* **1996**, *100*, 11375-11381.
- (30) Kibler, L. A.; El-Aziz, A. M.; Hoyer, R.; Kolb, D. M. *Angew. Chem. Int. Edit.* **2005**, *44*, 2080-2084.
- (31) Casado-Rivera, E.; Gal, Z.; Angelo, A. C. D.; Lind, C.; DiSalvo, F. J.; Abruna, H. D. *ChemPhysChem* **2003**, *4*, 193-199.
- (32) Roychowdhury, C.; Matsumoto, F.; Mutolo, P. F.; Abruna, H. D.; DiSalvo, F. J. *Chem. Mater.* **2005**, *17*, 5871-5876.
- (33) Blasini, D. R.; Rochefort, D.; Fachini, E.; Alden, L. R.; DiSalvo, F. J.; Cabrera, C. R.; Abruna, H. D. *Surf. Sci.* **2006**, *600*, 2670-2680.
- (34) Alden, L. R.; Han, D. K.; Matsumoto, F.; Abruña, H. D.; DiSalvo, F. J. *Chem. Mater.* **2006**, *18*, 5591-5596.
- (35) Matsumoto, F.; Roychowdhury, C.; DiSalvo, F. J.; Abruna, H. D. *J. Electrochem. Soc.* **2008**, *155*, B148-B154.
- (36) Clavilier, J.; Armand, D.; Sun, S. G.; Petit, M. *J. Electroanal. Chem.* **1986**, *205*, 267-277.
- (37) Korzeniewski, C.; Climent, V.; Feliu, J. M. In *Electroanalytical chemistry: A series of advances*; Bard, A. J., Zoski, C., Eds.; CRC Press: Boca Raton, 2012; Vol. 24, p 75-169.
- (38) Rodes, A.; Elachi, K.; Zamakhchari, M. A.; Clavilier, J. *J. Electroanal. Chem.* **1990**, *284*, 245-253.
- (39) Itaya, K.; Sugawara, S.; Sashikata, K.; Furuya, N. *J. Vac. Sci. Technol. A* **1990**, *8*, 515-519.
- (40) Gomez-Marin, A. M.; Feliu, J. M. *Electrochim. Acta* **2012**, *82*, 558-569.

- (41) Clavilier, J.; Feliu, J. M.; Aldaz, A. *J. Electroanal. Chem.* **1988**, *243*, 419-433.
- (42) Garcia-Araez, N.; Climent, V.; Feliu, J. In *Interfacial phenomena in electrocatalysis*; Vayenas, C. G., Ed.; Springer New York: 2011; Vol. 51, p 1-105.
- (43) Delley, B. *J. Chem. Phys.* **1990**, *92*, 508-517.
- (44) Delley, B. *Phys. Rev. B* **2002**, *66*, 155125.
- (45) Hammer, B.; Hansen, L. B.; Nørskov, J. K. *Phys. Rev. B* **1999**, *59*, 7413-7421.
- (46) Delley, B. *J. Chem. Phys.* **2000**, *113*, 7756-7764.
- (47) Delley, B. *Mol. Simul.* **2006**, *32*, 117-123.
- (48) Neugebauer, J.; Scheffler, M. *Phys. Rev. B* **1992**, *46*, 16067-16080.
- (49) Monkhorst, H. J.; Pack, J. D. *Phys. Rev. B* **1976**, *13*, 5188-5192.
- (50) Maciá, M. D.; Herrero, E.; Feliu, J. M.; Aldaz, A. *Electrochem. Commun.* **1999**, *1*, 87-89.
- (51) Maciá, M. D.; Herrero, E.; Feliu, J. M.; Aldaz, A. *J. Electroanal. Chem.* **2001**, *500*, 498-509.
- (52) Herrero, E.; Climent, V.; Feliu, J. M. *Electrochem. Commun.* **2000**, *2*, 636-640.
- (53) Leiva, E.; Iwasita, T.; Herrero, E.; Feliu, J. M. *Langmuir* **1997**, *13*, 6287-6293.
- (54) Kim, J.; Rhee, C. K. *Electrochem. Commun.* **2010**, *12*, 1731-1733.
- (55) Mulliken, R. S. *J. Chem. Phys.* **1955**, *23*, 1833-1840.
- (56) Hirshfeld, F. L. *Theor. Chim. Acta* **1977**, *44*, 129-138.
- (57) Peng, B.; Wang, H.-F.; Liu, Z.-P.; Cai, W.-B. *J. Phys. Chem. c* **2010**, *114*, 3102-3107.
- (58) Lin, W. F.; Sun, S. G.; Tian, Z. W. *J. Electroanal. Chem.* **1994**, *364*, 1-7.
- (59) Smoluchowski, R. *Phys. Rev.* **1941**, *60*, 661-674.
- (60) Wang, H.-F.; Liu, Z.-P. *J. Phys. Chem. C* **2009**, *113*, 17502-17508.
- (61) Gao, W.; Keith, J. A.; Anton, J.; Jacob, T. *J. Am. Chem. Soc.* **2010**, *132*, 18377-18385.
- (62) Neurock, M.; Janik, M.; Wieckowski, A. *Faraday Discuss.* **2009**, *140*, 363-378.
- (63) Perales-Rondón, J. V.; Ferre-Vilaplana, A.; Feliu, J. M.; Herrero, E. *J. Am. Chem. Soc.* **2014**, *136*, 13110-13113.
- (64) Perales-Rondón, J. V.; Herrero, E.; Feliu, J. M. *Electrochim. Acta* **2014**, *140*, 511-517.
- (65) Blais, S.; Jerkiewicz, G.; Herrero, E.; Feliu, J. M. *J. Electroanal. Chem.* **2002**, *519*, 111-122.

Table 1. DFT calculated adsorption energies of different adatoms deposited on the fcc and hcp sites of the Pt(111) surface.

	Adsorption energy (fcc) /eV	Adsorption energy (hcp) /eV
Pb	4.44	4.45
Bi	3.29	3.26
Sb	4.24	4.20
Te	3.68	3.55
As	4.40	4.33
Se	3.96	3.77
S	4.47	4.22
Pt	3.75	3.59

Table 2. Electronegativities and DFT calculated partial charges under two different formalisms for different adatoms deposited on a fcc site of the Pt(111) surface.

	Electronegativity (Pauling)	Adatom charge Mulliken/e	Adatom charge Hirshfeld/e
Pb	1.87	0.64	0.40
Bi	2.02	0.50	0.32
Sb	2.05	0.41	0.27
Te	2.10	0.28	0.20
As	2.18	0.18	0.11
Pt	2.28	0.05	0.08
Se	2.55	0.02	0.06
S	2.58	0.27	-0.02

Table 3. DFT calculated physisorption energies for a hydrated formic acid molecule on the adatom modified Pt(111) surface and the distance between the adatom and the O atom of the carbonyl group in the formic acid molecule.

	Distance/Å	Energy/eV
Pb	3.26	-0.27
Bi	3.07	-0.26
Sb	3.36	-0.21
Te	3.69	-0.21
As	3.82	-0.19

Table 4. DFT calculated activation energies for the cleavage of the C-H bond from the chemisorbed HCOO fragment on the adatom and the total energy of the process. Negligible activation energy is denoted as “dh”.

	Activation barrier/eV	Energy/eV
Pb	dh	- 1.03
Bi	dh	- 1.25
Sb	0.11	- 1.24
Te	dh	- 1.59
As	0.09	- 1.53

Table 5. DFT calculated adsorption energies for HCOO_m and HCOO_b on unmodified and Bi modified Pt(111) surfaces, using the adsorption energy of HCOO_b on the unmodified Pt(111) surface as reference

Species	$\Delta E/\text{eV}$
HCOO_m on Pt(111)	0.58
HCOO_b on Bi-Pt(111)	0.17
HCOO_m on Bi-Pt(111)	0.44

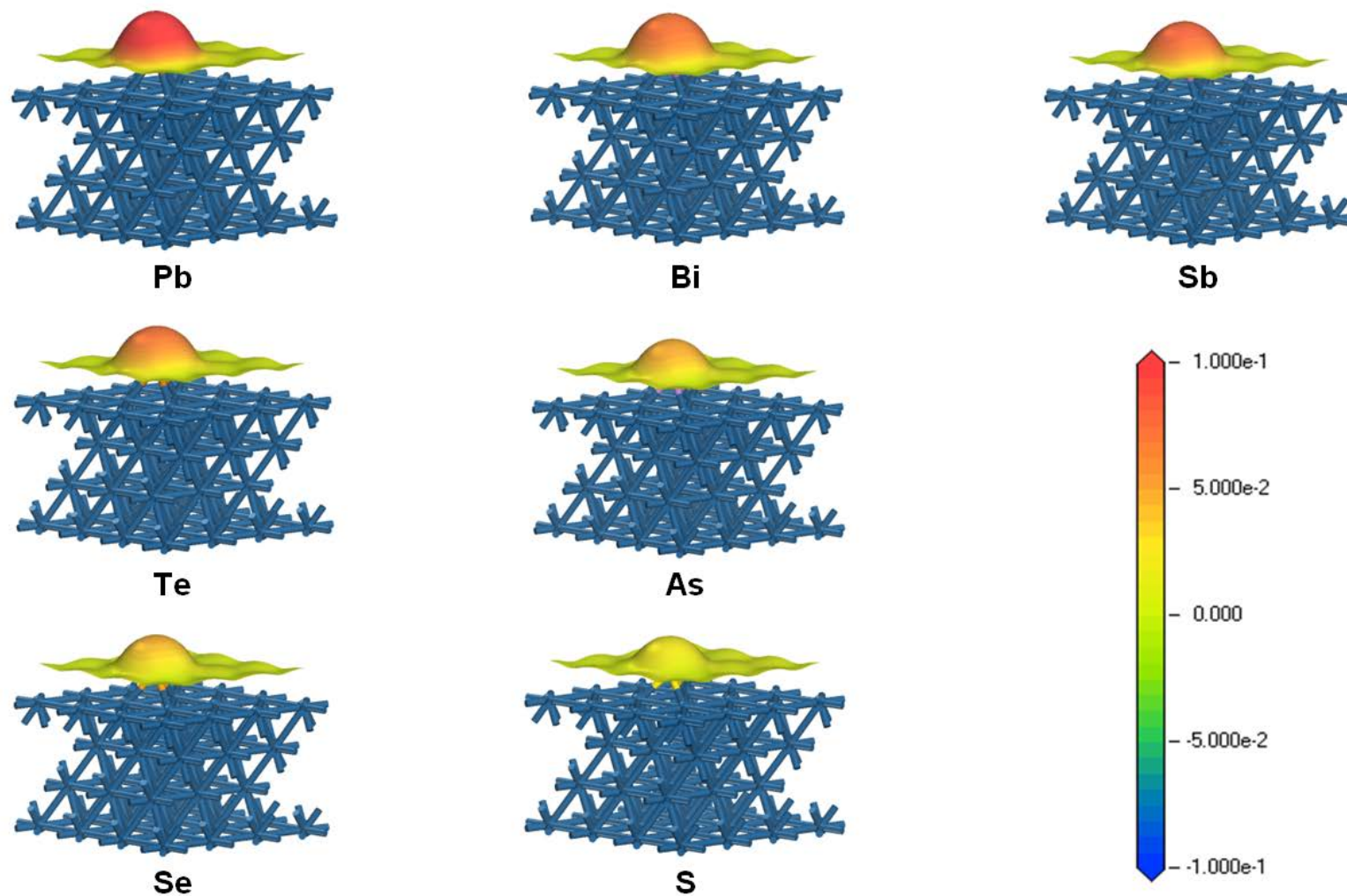


Figure 1. Electrostatic potential [Ha/e] mapped on electron iso-density surface for a density value $\rho = 0.01 \text{ e}/\text{\AA}^3$ for Pb, Bi, Sb, Te, As, Se and S adatoms adsorbed on the Pt(111) surface.

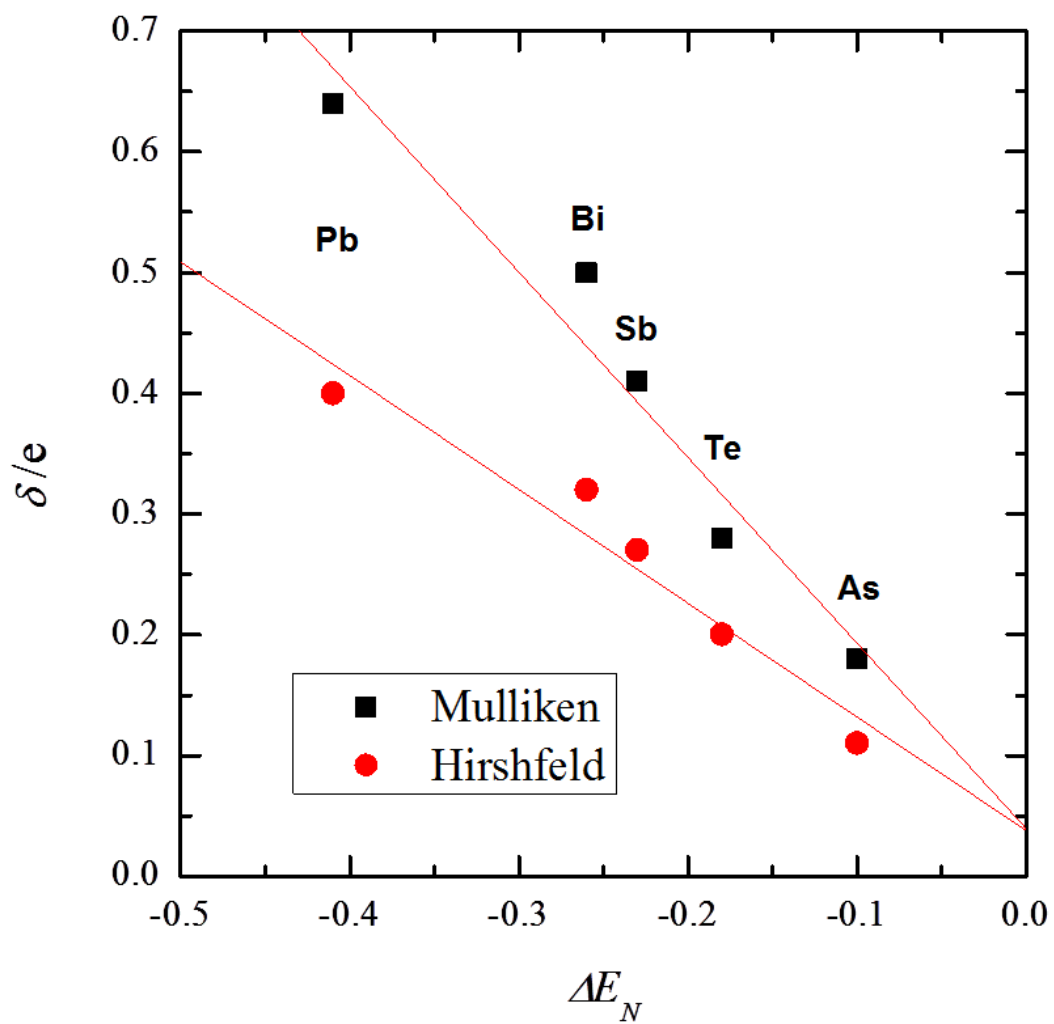


Figure 2. Partial charges (δ) vs. difference of electronegativity between the adatom and platinum (ΔE_N) estimated under two different formalisms.

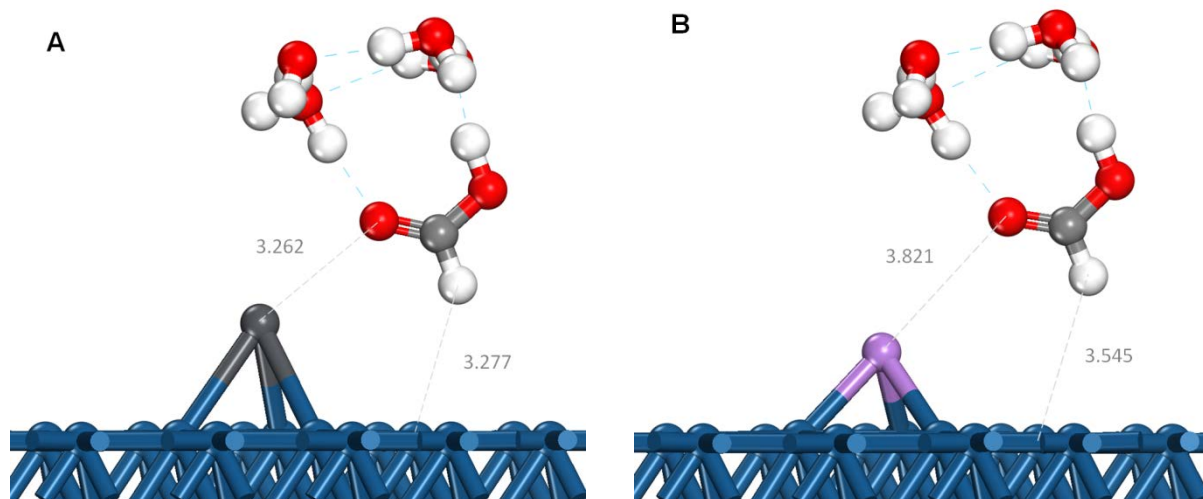


Figure 3. Physisorbed geometries of hydrated formic acid molecules on (A) Pb and (B) As modified Pt(111) surfaces.

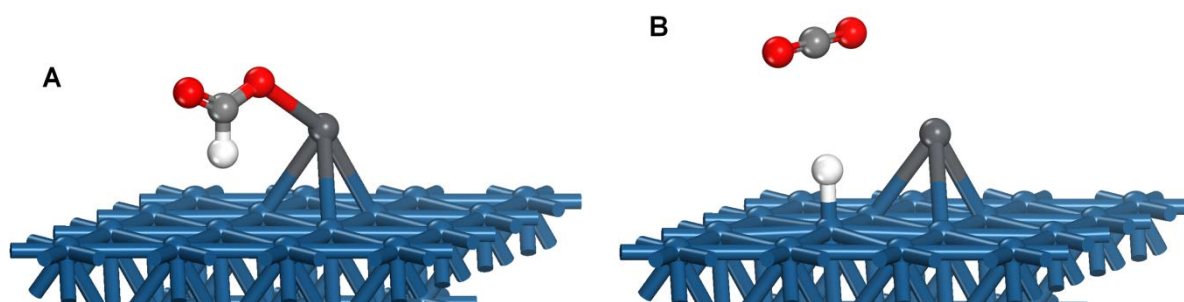


Figure 4. Geometries of A) the HCOO fragment chemisorbed on the Pb-Pt(111) surface and B) the final products yielded from the formic acid oxidation.

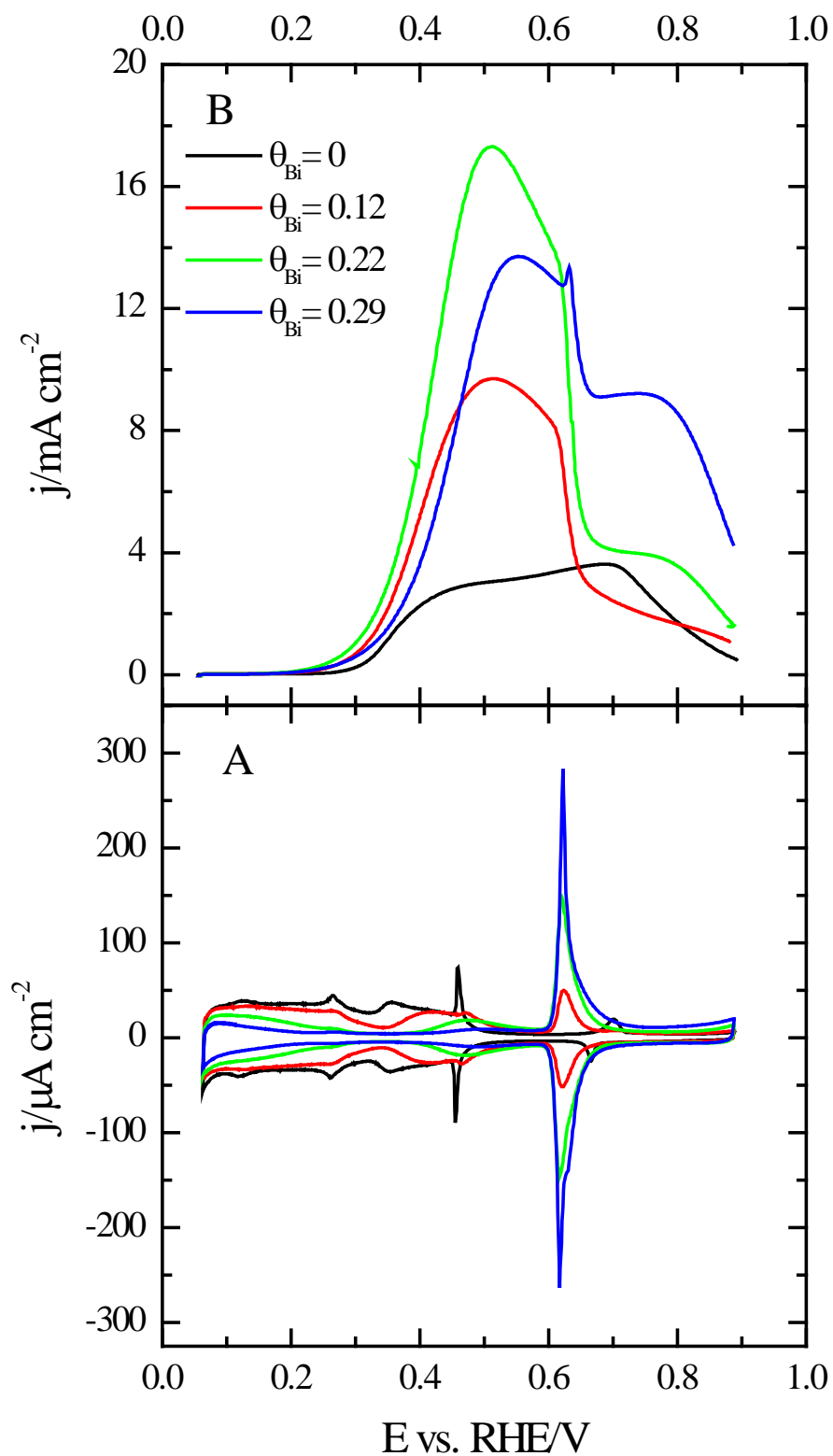


Figure 5. Voltammetric profiles for the Bi-Pt(111) electrode for different bismuth coverages in A) a 0.5 M H_2SO_4 solution and B) in a 0.5 M H_2SO_4 + 0.1 M HCOOH solution at 25 °C. Scan rate: 50 mV s^{-1} .

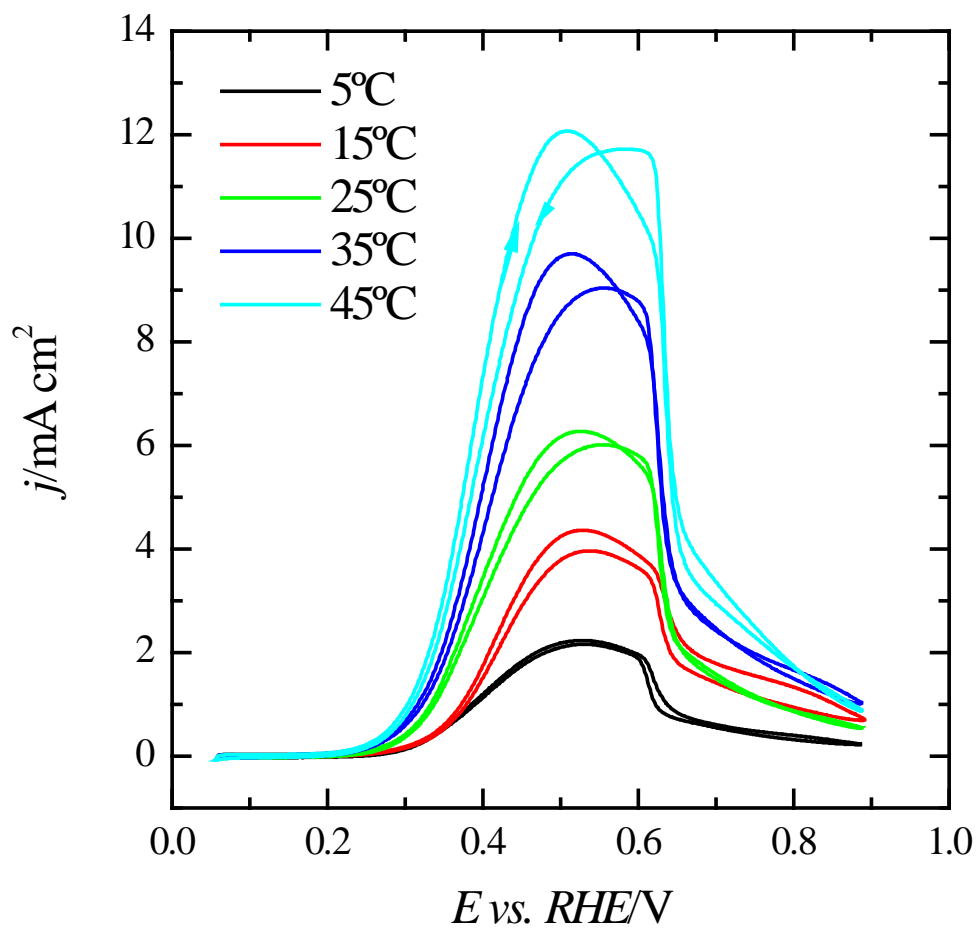


Figure 6. Voltammetric profiles for the Bi-Pt(111) electrode with $\theta_{\text{Bi}}=0.12$ in a 0.5 M H_2SO_4 + 0.1 M HCOOH solution at different temperatures. Scan rate: 50 mV s^{-1} .

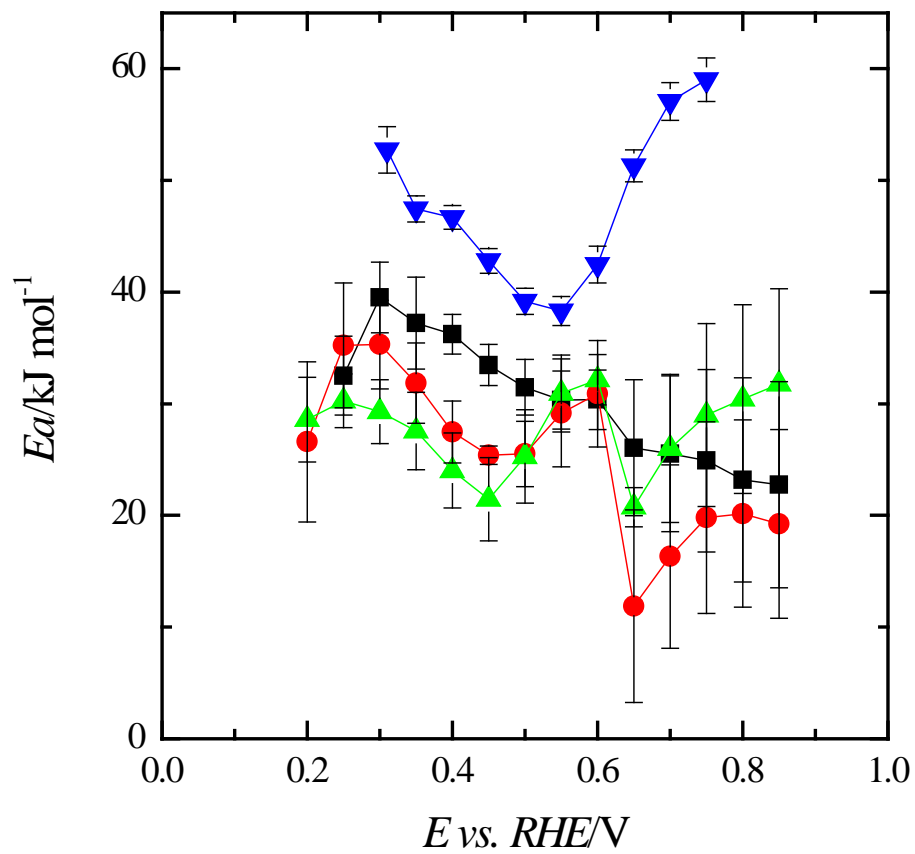


Figure 7. Measured activation energies at different potentials for different Bi coverages on the Pt(111) electrode (\blacktriangledown) $\theta_{Bi}=0.00$; (\blacksquare) $\theta_{Bi}=0.12$ (\bullet) $\theta_{Bi}=0.22$ and (\blacktriangle) $\theta_{Bi}=0.28$.

TOC

