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Additional Information

# Use of ancient copper slags in Portland cement and alkali activated cement matrices

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## Abstract

Some Chilean copper slag dumps from the nineteenth century still remain, without a proposed use that encourages recycling and reduces environmental impact. In this paper, the copper slag abandoned in landfills is proposed as a new building material. The slags studied were taken from Playa Negra and Púquios dumps, both located in the region of Atacama in northern Chile. Pozzolanic activity in lime and Portland cement systems, as well as the alkali activation in pastes with copper slag cured at different temperatures, was studied. The reactivity of the slag was measured using thermogravimetric analysis (TGA), scanning electron microscopy (SEM), X-ray diffraction (XRD), electrical conductivity and pH in aqueous suspension and Fourier Transform Infrared Spectroscopy (FTIR). Furthermore, copper slag-Portland cement mortars with the substitution of 25% (by weight) of cement by copper slag and alkali-activated slag mortars cured at 20 and 65°C were made, to determine the compressive strength. The results indicate that the ancient copper slags studied have interesting binding properties for the construction sector.

**Keywords:** Copper slag; pozzolanic activity; alkali activation; geopolymer.

## 1. Introduction

Industrial slags are usually classified as ferrous and nonferrous slags, and may be used in the manufacture of new materials. Slags from extracting metal processes are conventionally considered as waste materials (Nazari and Sanjayan, 2015). In construction, copper slag (CS) is used in fillings, embankments, drainage, ballast, as aggregates in mortars, concrete (Thomas and Gupta, 2013; Mithun and Narasimhan, 2015) and asphalt mixtures (Hassan and Al-Jabri, 2011), as abrasive material (Kambham et al., 2007), as a cement substitute (Shi and Qian, 2000; Moura and Coutinho, 2004; Zain et al., 2004; Al-Jabri et al., 2006; Moura et al., 2007; Taha et al., 2007; Sánchez de Rojas et al., 2008; Brindha et al., 2010; Chew and Bharati, 2010; Brindha and Nagan, 2011; Najimi et al., 2011; Peyronnard and Benzaazoua, 2011), as raw material in the manufacture of Portland cement (PC) (Ariño and Mobasher, 1999; Chockalingam et al., 2013;

García Medina et al., 2006; Gorai et al., 2003; Kosmatka et al., 2003) and also as binders in the manufacture of alkali-activated (AA) cements (Deja and Malolepszy, 1989, 1994; Shi et al., 2008; Iacobescu et al., 2013).

Regarding the use of CS aggregates, we can report that there are a few technical specifications (Federal Highway Administration Research and Technology. U.S. Department of Transportation, n.d.) and two standards governing their use as aggregates for concrete: one in Korea (Korean Standards Association, 2000) and one in Japan (Japan Mining Industry Association, 2003). The residue from the flotation tailings of CS has also been used in cement admixtures and cement manufacturing (Alp et al., 2008; Onuaguluchi, 2012).

According to the method of cooling employed, CS can be grouped into two types: air-cooled slags, which are black glass with low water absorption, and water-cooled slags which are amorphous, granulated and present greater water absorption capacity (Onuaguluchi, 2012). CS generated nowadays is composed "Fe (30–40%), SiO<sub>2</sub> (35–40%), Al<sub>2</sub>O<sub>3</sub> (≤10%), CaO (≤10%) and Cu (0.5–2.1%)" (Gorai et al., 2003).

In Chile, there are CS dumps that were produced in the nineteenth century that have environmental effects on the soil (destruction of beaches and shoreline pollution) and landscape (visual pollution barriers in natural channels), as well as social (alteration of Historic Places) and economic (land use) impacts. Chilean law classifies these slags as non-hazardous waste (Ministerio de Salud de Chile, 2004). Often, "the statistical analysis provides a basis for predicting with high probability that the composition of the CS and their leachates will remain in the non-hazardous range" (Alter, 2005). Moreover, they can degrade over time, and may potentially be a source of toxic metals (Piatak et al., 2004). So far there has been no proposal to use former CS landfills and reduce these effects, or to use them as raw materials in the manufacture of mortars and concretes based on PC and/or AA cement.

The aim of this study is to determine the feasibility of using two CS produced in the nineteenth century, Playa Negra and Púquios, which are abandoned landfills located in the region of Atacama in Chile, as building materials. Pozzolanic activity in lime and Portland cement systems, as well as the AA in pastes with copper slag cured at different temperatures, was studied. Moreover, CS-PC mortars with the substitution of 25% (by weight) of cement by CS and AA slag mortars cured at 20 and 65°C were made, to determine the compressive strength.

## 2. Experimental

### 2.1. Materials

CS samples of the nineteenth century, originating from the landfills Playa Negra (CS\_PN) and Púquios (CS\_PQ), both located in the region of Atacama in northern Chile, were used. The CS\_PQ is located 150 km away from CS\_PN. The PC used in cement pastes and mortars, is

CEM I-52.5 R type according to EN 197-1 (UNE-EN 197-1:2011, 2011). The chemical composition of the slags and the PC used are shown in Table 1. In mortars CS/PC, standard sand (Normensand, Germany) was used, while in the AA mortars silica sand (Caolines Lapedra, Valencia, Spain) was used, according to EN 196-1 standard (UNE-EN 196-1:2005, 2005). As an activator for the AA reaction, 98% pure sodium hydroxide in pellet form (Panreac) and sodium silicate solution (SiO<sub>2</sub>=28%, Na<sub>2</sub>O=8%, H<sub>2</sub>O=64%, by weight, Merck) were used. In the conductimetric method (Tashima et al., 2014), for the testing of pozzolanic reactivity, an aqueous suspension of calcium hydroxide powder was used (minimum 95% purity, Panreac). In the manufacture of cement mortars and pastes, water from the Valencia city network and deionised water were used.

Table 1. Chemical composition of PC and CS (Wt. %) and Density (g/cm<sup>3</sup>).

	SiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	Fe <sub>2</sub> O <sub>3</sub>	CaO	MgO	SO <sub>3</sub>	K <sub>2</sub> O	Na <sub>2</sub> O	Other	LOI	Density
PC	20.80	4.60	4.80	65.60	1.20	1.70	1.00	0.07	0.23	2.02	3.07
CS_PN	39.14	7.76	30.48	13.41	2.09	0.46	1.50	1.05	3.60	-2.11*	3.16
CS_PQ	38.33	8.17	20.40	26.10	2.14	0.26	0.78	0.64	3.59	-1.50*	2.96

\* mass gain probably due to oxidation FeO and S<sup>2-</sup> (sulphides) (Sánchez de Rojas et al., 2008).

The CS was first milled using a jaw crusher; after this, a ball mill with alumina balls that were 18 mm in diameter was used. Particle size distribution was determined using a laser diffraction granulometer (Mastersizer 2000, Malvern Instruments). The final average diameter for both milled CS samples was 15 µm and 90% of particles were smaller than 40 µm after 35 minutes grinding time. The particle size distribution curves of both ground forms of CS are shown in Fig. 1.

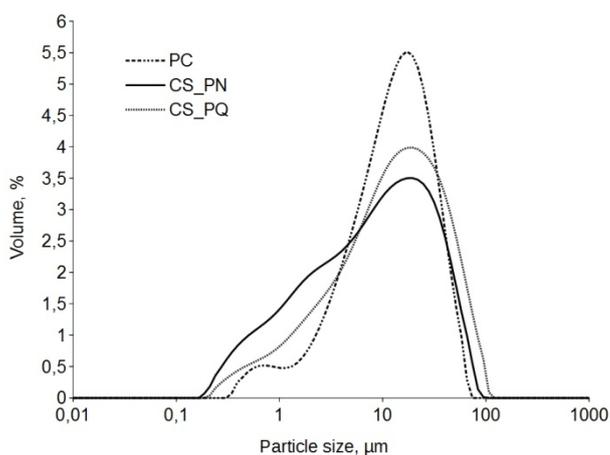


Fig. 1. Particle size distribution curve of PC and ground CS.

The CS chemical compositions were determined by X-ray fluorescence (XRF MAGIC PRO Philips, mod. PW2400) and are shown in Table 1. The results show that the two samples of CS contain a very significant amount of iron oxides and silica. The silica content of both samples is

very similar; however, the iron oxide content is higher in the CS\_PN. They also have a high content of CaO, which may indicate their feasibility as a precursor material for alkali activation (Canfield et al., 2014). Moreover, it can be seen that the CaO content in CS\_PQ is twice the content in CS\_PN. **The high content of CaO and Fe<sub>2</sub>O<sub>3</sub> less, with respect to copper slag used in other studies (Gorai et al., 2003) is due to the contribution of fluxes such as silica and lime used during smelting. It is usual to add silica to capture separating copper iron during smelting, and lime gives flow properties within the furnace slag.**

The X-ray diffraction (XRD, Brucker AXS D8 Advance) patterns were obtained using a Cu-K $\alpha$  radiation under conditions of 40 kV and 20 mA, in the range  $2\theta=5-70^\circ$ . Fig. 2 shows that CS\_PQ is fully amorphous; this is evidenced by the strong deviation of the baseline between 20 and 40° 2 $\theta$  values. Only trace amounts of diopside type crystalline compounds, fayalite and clinoferrosilite were detected. Furthermore, the mineralogical CS\_PN showed a structure that mainly consisted of diopside, magnetite and fayalite; clinoferrosilite was present to a lesser extent.

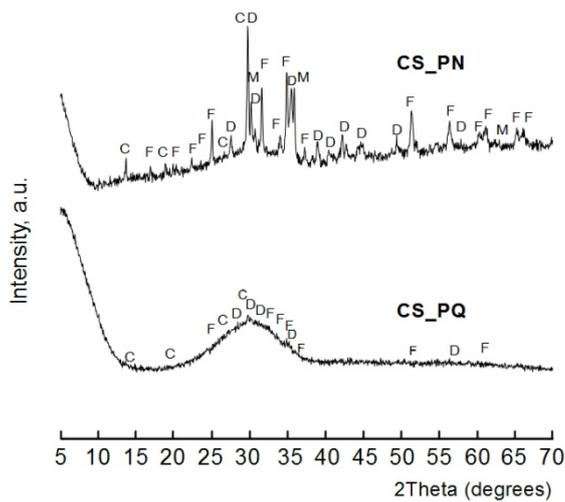


Fig. 2. X-ray diffractogram patterns of raw CS\_PN and CS\_PQ materials. D: Diopside (CaMgSiO<sub>3</sub>; PDFcard 190239); F: Fayalite (Fe<sub>2</sub>SiO<sub>4</sub>; PDFcard 340178); C: Clinoferrosilite (FeSiO<sub>3</sub>; PDFcard 170548); M: Magnetite (Fe<sub>3</sub>O<sub>4</sub>; PDFcard 190629).

Thermogravimetric analysis (TGA, Mettler-Toledo 850) of CS samples was carried out at ranges of **35-1000°C (alumina crucible), with the heating rate being 20°C/min** in air atmosphere. TG and DTG curves of CS\_PN and CS\_PQ samples are plotted in Fig. 3. In TGA curves, CS\_PN and CS\_PQ samples gain masses of 3.15 and 2.15% respectively, probably due to the oxidation of **Fe (II) compounds** and sulphides S<sup>2-</sup> (Sánchez de Rojas et al., 2008). In the **CS samples part** of the mass gain could be due to fayalite oxidation according to the equation  $\text{Fe}_2\text{SiO}_4 + \frac{1}{2}\text{O}_2 \rightarrow \alpha\text{Fe}_2\text{O}_3 + \text{SiO}_2$  (Gorai et al., 2003).

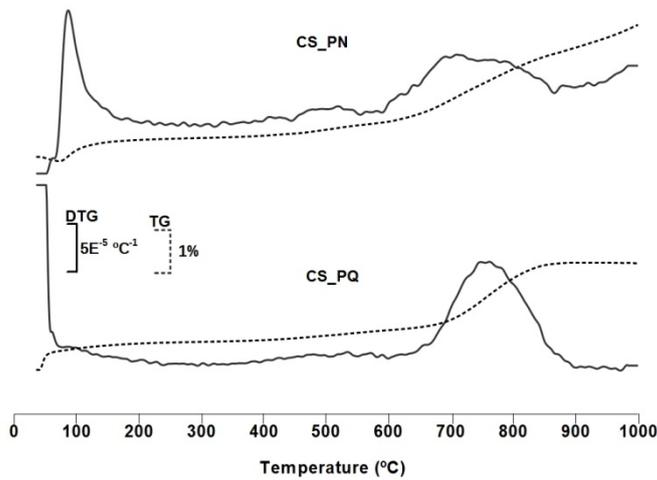


Fig. 3. TG and DTG curves for CS\_PN and CS\_PQ.

The samples were also evaluated by Fourier Transform Infrared Spectroscopy (FTIR Bruker Tensor 27 MIR, range of  $7800\text{-}370\text{ cm}^{-1}$ ). The spectrum for CS PQ of Fig. 4 shows peaks in the range  $3750\text{-}3600\text{ cm}^{-1}$ , which are associated with OH valence vibrations, both symmetric and asymmetric (Darder et al., 2014). Furthermore, for both samples, the peaks in the range  $1000\text{-}850\text{ cm}^{-1}$  can be attributed to Si-O asymmetric stretching vibrations of the  $\nu_3$  mode in  $\text{SiO}_4$  (Mihailova and Mehandjiev, 2010). In the range  $450\text{-}500\text{ cm}^{-1}$ , vibrations associated to  $\nu_4(\text{O-Si-O})$  deformation can be observed (Fernández-Jiménez, 2000). In general, the FTIR spectra show that both slags exhibit vibrations primarily associated with Si-O bonds.

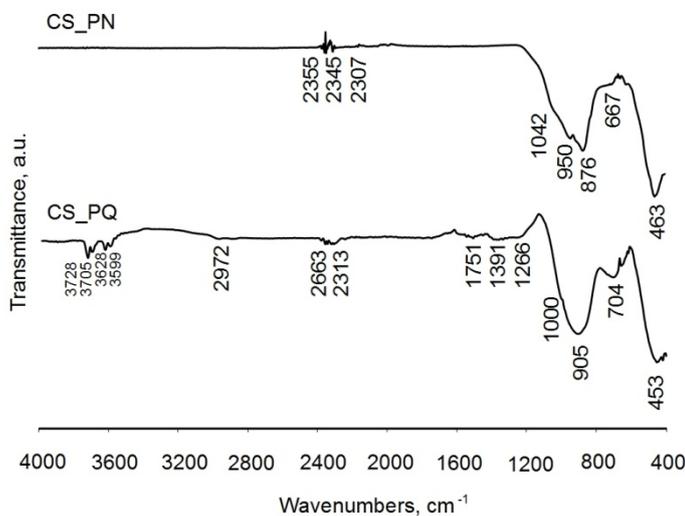


Fig. 4. FTIR spectra of CS\_PQ and CS\_PN samples.

Microscopic analysis and morphology of CS samples were studied using scanning electron microscopy (SEM, JEOL-JSM-6300) equipped with energy dispersive X-ray spectroscopy (EDX) and using a voltage of 20 kV for imaging. Fig. 5 shows that both ground CS samples presented crushed particles with angular shapes.

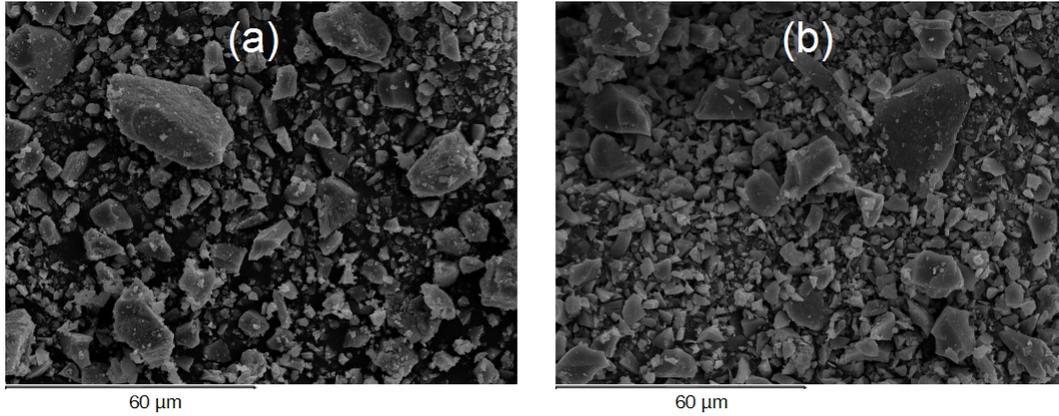


Fig. 5. SEM micrographs of ground CS: a) CS\_PN; b) CS\_PQ.

## 2.2. Procedures

### 2.2.1. Copper slag alkali-activated cement pastes

CS AA cement pastes were prepared by mixing CS with an alkaline activator solution (by weight). Mixes are coded as '25/5/1.45', where 25 is the amount of water per 100 g of CS, 5 is the molality of  $\text{Na}^+$  in the activating solution and 1.45 is the  $\text{SiO}_2/\text{Na}_2\text{O}$  molar ratio in the activating solution. As shown in Table 2 a set of pastes samples were cured at 20°C and 95% RH for 7 and 28 days and other set of pastes at 65°C in thermostatically controlled bath for 1, 3 and 7 days. The alkaline activator solution was prepared by dissolving NaOH pellets in water and adding the required amount of sodium silicate solution. Paste samples were milled, washed with acetone to stop the reaction process, and stored sealed until characterisation. TGA studies on AA pastes were carried out using an sealed aluminum crucibles with a pinhole lid, from 35-600°C at 10°C/min heating rate in air atmosphere. Microscopic analysis and morphology of AA pastes were studied using SEM with a voltage of 20 kV for imaging.

Table 2. Alkali-activated cement pastes nomenclature.

Sample	Curing time, days	Curing temperature, °C
CS_PN AA20	7 and 28	20
CS_PN AA65	1, 3 and 7	65
CS_PQ AA20	7 and 28	20
CS_PQ AA65	1, 3 and 7	65

### 2.2.2. Portland cement mortars containing CS

Mortar cubes of 40×40×40 mm replacing 25% (in weight) of PC by CS (PC/CS\_PN and PC/CS\_PQ mortars) were manufactured and tested in compression at ages of 28 and 90 days. PC mortars were also made as control samples. In all mortars, the water/binder ratio (w/b) was 0.44 (the binder being the sum of PC and CS). The mixing procedure was carried out according UNE-EN 196-1 (UNE-EN 196-1:2005, 2005). Mortars were cured at 20°C under a saturated solution of calcium hydroxide. The mortar compositions are summarised in Table 3.

Table 3. Mortar composition nomenclature.

	<b>PC</b>	<b>CS</b>	<b>Sand</b>	<b>w/b</b>
	<u>g</u>	<u>g</u>	<u>g</u>	
<b>Mortar control</b>	450.0	0	1350	0.44
<b>Mortar PC/CS_PN</b>	337.5	112.5	1350	0.44
<b>Mortar PC/CS_PQ</b>	337.5	112.5	1350	0.44

### 2.2.3. Copper slag alkali-activated cement mortars

AA cement mortar cubes with a size of 40×40×40 mm were manufactured. A set of mortars samples was cured in a chamber at 20°C for 7, 28 and 90 days and other set of mortars were cured in a thermostatically controlled bath at 65°C for 3, 7 and 28 days. For each samples set, 450 grams of copper slag and 1350 grams of sand were used. The molar ratio of the alkaline activators  $\text{SiO}_2/\text{Na}_2\text{O}$  was 1.45 and the w/CS was 0.30. For analysis of the results of mortar strength, the approach suggested in UNE 83485:1996 EX (UNE 83485:1996 EX, 1996) was applied.

### 2.2.4. Conductimetric method for pozzolanic reactivity evaluation

To assess the pozzolanic reactivity of the slag, the Tashima method was used (Tashima et al., 2014). This method consists of measuring the pH and electrical conductivity evolution for 7 days of different mixtures of calcium hydroxide-saturated aqueous suspensions and pozzolan (CH:pozzolan in mass ratio, one gram total in 50 ml of desionized water for each ratio) at different temperatures: 40, 50 and 60°C.

A pH meter Crison micro PH2001 and an electrical conductivity meter Crison micro CM2201 were used to measure the pozzolanic reaction progress. A shaking water bath JULABO– SW22 working in the temperature range of 20–99.9°C was used to control the test temperature. At the beginning of the test, all of the prepared suspensions were saturated in calcium hydroxide with an excess of calcium hydroxide in the solid state. During the testing period, a reduction in the pH and electrical conductivity values can be observed for some of the suspensions, indicating the unsaturated state of the suspension with respect to calcium hydroxide due to the pozzolanic reaction progress. The reductions in pH and electrical conductivity values are due to the reaction of dissolved  $\text{Ca}^{+2}$  and  $\text{OH}^-$  ions with pozzolan particles to form stable and insoluble products (Luxán et al., 1989). The presence of calcium hydroxide in the solid state maintained the concentration of dissolved  $\text{Ca}(\text{OH})_2$  at an approximately constant level in the suspension (saturated system) and, when the ions dissolved in the solution are consumed, solubilisation of calcium hydroxide took place quickly, again reaching saturation of the solution. The solubilisation process is finished when calcium hydroxide in solid state is exhausted. From this moment (unsaturated system), when the CH in solid state is totally dissolved, a reduction on pH and electrical conductivity values can be detected, indicating the beginning of the unsaturation step. In this study, different CH:CS mixes,

1:9, 2:8, 3:7, 4:6 and 5:5 (mass ratio, one gram total in 50 ml of desionized water for each ratio) for 40, 50 and 60°C, were tested.

### 3. Results and discussion

#### 3.1. Alkali-activated copper slag cement pastes

In Fig. 6, XRD patterns for AA pastes cured at 20°C for 28 days are shown. The XRD pattern for CS\_PN AA paste (Fig. 6a) showed the same peaks as those observed for the original slag. This means that the crystalline products present in the original slag did not evolve during the alkali activation reaction. Additionally, no evidence of the zeolitic phase formation was found. The reaction products formed were amorphous. CS\_PQ slag did not show any crystallized products and after alkali activation no changes were noticed (Fig. 6b); again, the reaction compounds were amorphous and there were no zeolitic phases. AA pastes cured at 65°C for 7 curing days yielded very similar XRD patterns to those obtained at 20°C for 28 days: this means that the curing temperature in the 20-65°C range did not significantly affect the reaction products.

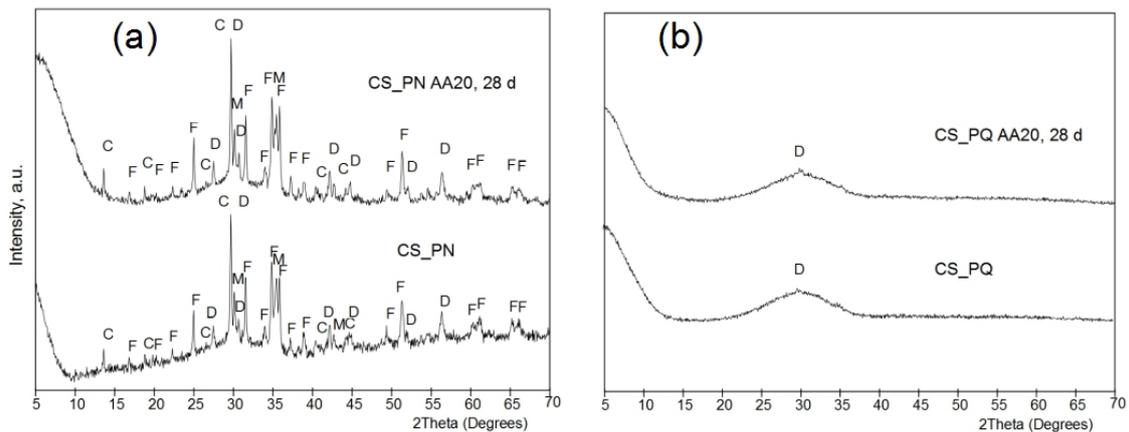


Fig. 6. X-ray diffractograms of CS and AA pastes after curing at 20°C for 28 days. (a) CS\_PN and (b) CS\_PQ.

FTIR analysis of the AA pastes is shown in Fig. 7. In general, the presence of bands in the range 3750-3600  $\text{cm}^{-1}$  can be observed, attributed to valence vibrations OH, asymmetric and symmetric (Darder et al., 2014); the presence of OH groups is related to the hydration products. The band at the 1450-1380  $\text{cm}^{-1}$  is attributed to O-C-O deformation vibration in carbonates due to carbonation products (may be an interaction of air and alkali hydroxide took place). Bands in the range 1000-850  $\text{cm}^{-1}$  are attributed to the polymerisation process of  $\text{SiO}_4$ - linking tetrahedra. For CS\_PN slag, the peak at 876  $\text{cm}^{-1}$  was maintained during the reaction progress (both at 20 and 65°C), and was attributed to crystalline matter (fayalite, diopside, clinoferrite); however, an increase in intensity for the peak at 950  $\text{cm}^{-1}$  was produced: this behaviour was due to the formation of new Si-O-Si bonds from the reaction progress. For CS\_PQ slag, the broad peak

centred at  $905\text{ cm}^{-1}$  evolved with alkali activation; thus, a higher energy peak (about  $950\text{-}960\text{ cm}^{-1}$ ) was produced for cured materials. Bands in the range  $445\text{-}460\text{ cm}^{-1}$  were attributed to bending vibrations of Si-O/Al-O (Nath and Kumar, 2013; Komnitsas et al., 2013) which were present for all cured materials.

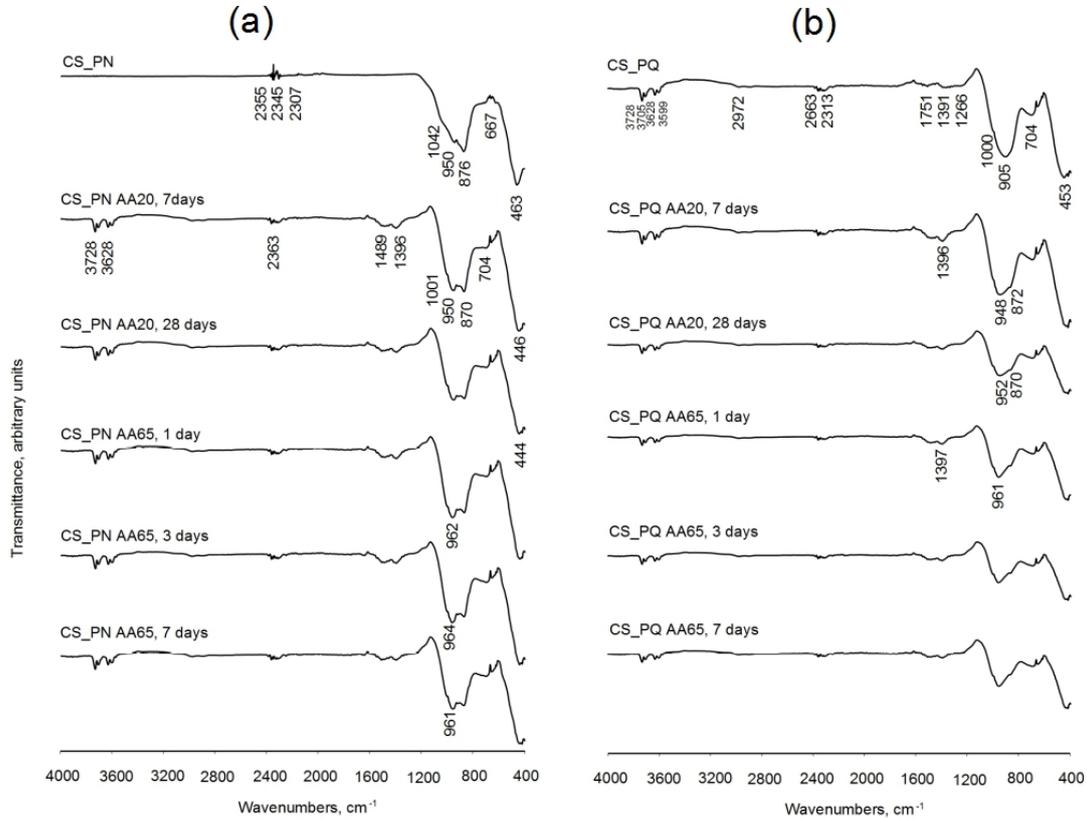


Fig. 7. (a) FTIR CS\_PN and (b) FTIR CS\_PQ, alkali activated slags cement pastes.

Table 4 shows thermogravimetric analysis data, summarising the total weight loss (TG,  $850\text{ Mettler-Toledo}$  from  $35\text{-}600^\circ\text{C}$  range, using sealed pin holed aluminium crucibles at a heating rate of  $10^\circ\text{C min}^{-1}$ ) of AA pastes cured at  $65^\circ\text{C}$  and  $20^\circ\text{C}$ . As observed in Fig. 8, all paste samples showed a DTG peak in the  $129\text{-}148^\circ\text{C}$  range, typical of AA gels. This loss of weight is attributed to chemically bond water and hydroxyl groups present in these samples (Tashima et al., 2012). In general, an increase in the total weight loss of pastes occurred when curing times increase for  $65$  and  $20^\circ\text{C}$ . This is evidence of progress of the AA process. When both slags are compared, slightly further progress of the reaction in the CS\_PQ AA, was generally observed, although the differences were more pronounced in the experiments carried out at  $65^\circ\text{C}$ .

Table 4. Thermogravimetric total loss of weight (%) in AA CS pastes cured at  $65$  and  $20^\circ\text{C}$ .

Curing temperature	$65^\circ\text{C}$			$20^\circ\text{C}$	
Cure time	1 day	3 days	7 days	7 days	28 days
CS_PN	6.01	8.27	7.81	4.67	6.25
CS_PQ	7.67	8.84	9.19	4.03	6.45

SEM micrographs of AA CS cements are shown in Fig. 9. It can be seen that the CS-reacted product created a matrix with a compacted appearance. However, it appears to be less porous when cured at a temperature of 65°C. When both slags are compared, lower porosity can be observed in CS\_PQ for both curing temperatures. This corroborates the reactivity of CS, as evidenced in thermogravimetry data shown in Table 4 and compressive strength development in Fig. 11.

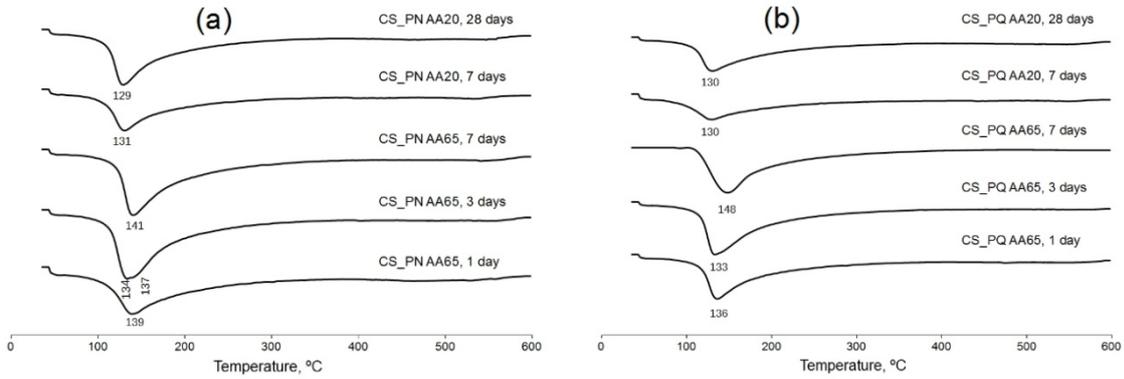


Fig. 8. Differential thermogravimetric curves for alkali-activated CS pastes (a) CS\_PN; (b) CS\_PQ.

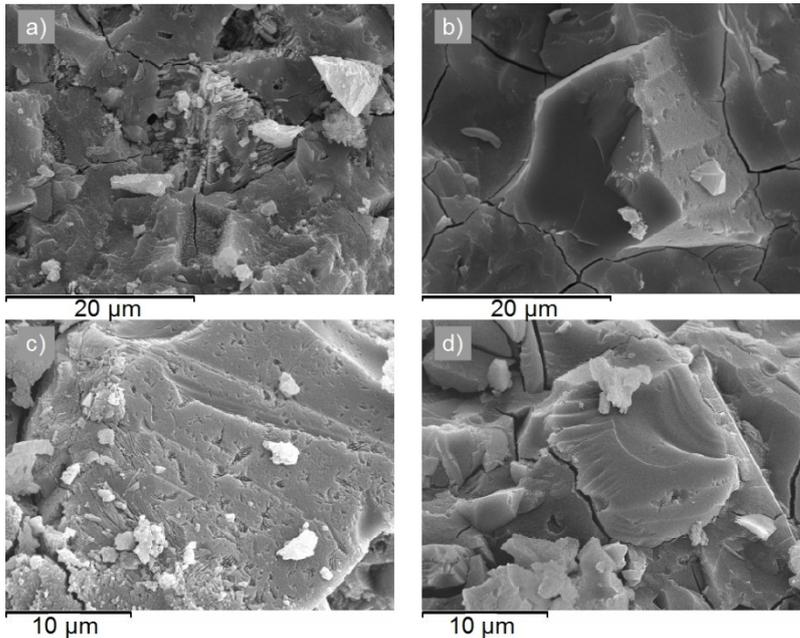


Fig. 9. Scanning electron microscope images of CS: a) CS\_PN AA65 cured for 7 days; b) CS\_PQ AA65 cured for 7 days; c) CS\_PN AA20 cured for 28; and d) CS\_PQ AA20 cured for 28 days.

### 3.2 Mortars

#### 3.2.1. Copper slag in Portland cement mortars

Compressive strengths for control mortars and mortar specimens with 25% replacement of PC by CS\_PN or CS\_PQ are shown in Fig. 10. Two control mortars were used, one for each

slag mortar, because they were prepared on different dates. In general, control mortar strength was higher than CS-containing mortars. The strength activity ratio (SAR) satisfactorily complied with the requirements of the European standard EN 450-1:2013 (*UNE-EN 450-1:2013, 2013*) for fly ashes. This standard requires an SAR of 75% for 28 days curing time and 85% for the 90 days of curing time. **CS\_PQ greater reactivity in the 90 days curing is observed.**

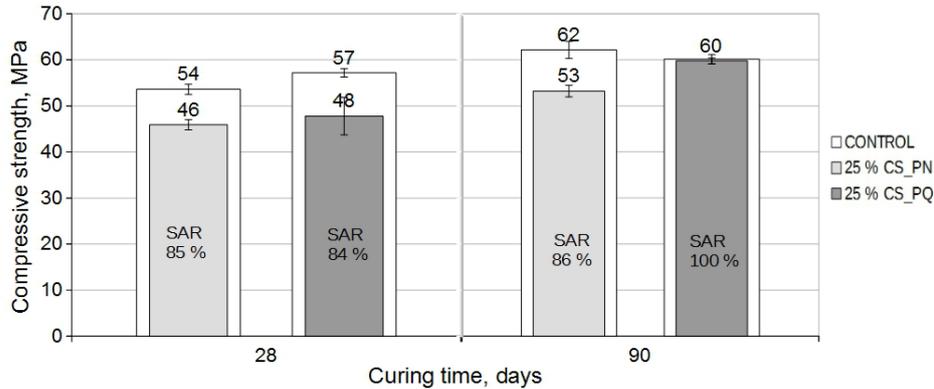


Fig. 10. Compressive strength of cement mortars and SAR. Control mortar and CS containing mortars with the replacement of 25% at curing ages of 28 and 90 days.

For an approach the cementing efficiency (*k*-factor) of the CS in mortars replacing a PC with 25% CS for ages 28 and 90 days it has been used Bolomey's model. The *k*-factor represents the amount of cement that can be replaced in a paste or mortar pozzolan to obtain the same value of mechanical strength and can be calculated using Eq. 1.

$$k = \frac{\left(\frac{\sigma}{a} - b\right) w - c}{p} \quad (\text{Eq. 1})$$

Where  $\sigma$  is the compressive strength, *a* and *b* are two constant calculated from compressive strength values for mortars with different *w/c* ratios and *p* is the amount of added pozzolan. The *k*-factors are shown in Table 5.

Table 5. *k*-factors, *a* and *b* for mortars containing the replacement of 25% of CS.

Sample	28 days			90 days		
	<i>a</i>	<i>b</i>	<i>k</i>	<i>a</i>	<i>b</i>	<i>k</i>
25% CS_PN	23.29	0.20	0.19	25.87	0.18	0.36
25% CS_PQ	23.29	0.20	0.35	25.87	0.18	0.85

### 3.2.2. Copper slag alkali-activated cement mortars.

The results of compressive strength test on AA CS mortars are shown in Fig. 11. The results of AA pastes are in agreement with those obtained for the compressive strength of AA mortars, wherein the compressive strength of CS\_PQ mortars was substantially higher than those made with CS\_PN (see Fig. 11). Data were obtained after 7, 28 and 90 days curing temperature of 20°C, and after 3 and 7 days curing at 65°C. An increase in compressive strength with curing

time can be observed for both slags and curing temperatures. In all cases, CS\_PQ mortars presented higher compressive strength than CS\_PN mortars, although this difference decreased when curing temperature or curing time increased. **CS\_PQ AA mortar shown a higher strength development at early ages and CS\_PN shown a better strength development for longer curing age or for higher curing temperature.** In the future, it would be interesting to design experiments with longer curing times to check whether compressive strengths of CS\_PQ and CS\_PN tend to equalize.

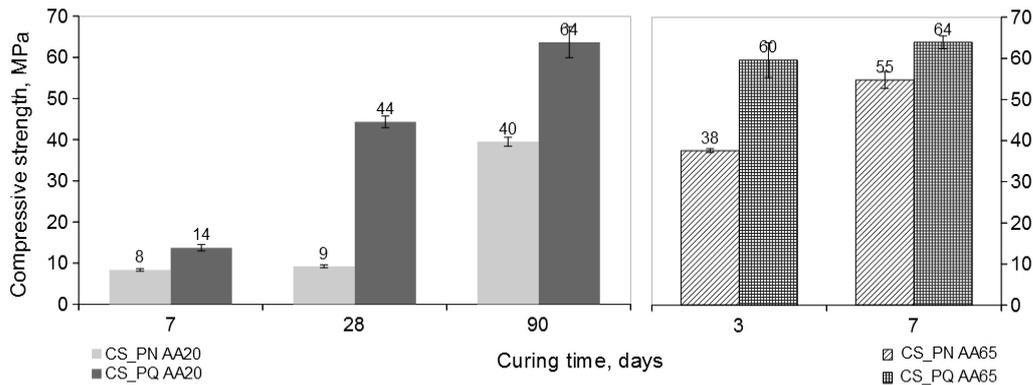


Fig. 11. Influence of curing time and curing temperature on compressive strength of **CS AA** mortars.

Higher compressive strength in CS\_PQ mortars can be explained by taking into account the higher reactivity of CS\_PQ, as evidenced in **SEM** studies. TG total loss of weight percentage for CS\_PQ **AA65** paste was **slightly** higher than CS\_PN **AA65** paste (see Table 4 and Fig. 7). The microstructure of CS\_PQ **AA** paste is less porous for both curing temperatures (see Fig. 9). The greatest CS\_PQ **AA** reactivity is probably related to the **more** amorphous character (see Fig. 2) which promotes the dissolution of the silica and alumina in the alkaline environment favouring **AA** reaction. It should be highlighted that mortars made with CS\_PQ **AA** reached a compressive strength of 64 MPa at 90 days of curing at 20°C. Similar compressive strengths for mortars cured for 3 days at 65°C and mortars cured for 90 days at 20 °C temperature were obtained.

### 3.3. Conductimetric method for pozzolanic reactivity evaluation

At 40 and 50°C results in loss of conductivity in both slags are studied below 30%, except the 1:9 and 2:8 mixtures to 60 °C. Therefore of no interest for comparison of materials.

At 60°C, the 1:9 suspension showed that the maximum electrical loss conductivity (LC) at 168 hours of testing was 83% and 40%, respectively, for CS\_PN and CS\_PQ. However, testing temperatures of 40 and 50°C deliver results below the line loss of 30% conductivity. Moreover, the variation in the pH of the 1:9 and 2:8 suspensions was not significant. Fig. 12 shows the variation of LC 1:9 and 2:8 suspensions at 60°C. According to Tashima et al. (2014), the unsaturation time that is associated with the  $LC(\%)=30\%$  could be used to assess the pozzolanic reactivity. Tashima et al. (2014) proposed three different zones, as shown Fig. 13: Zone 1 for low reactive pozzolans, Zone 2 for medium reactivity of the pozzolans and Zone 3 for highly reactive pozzolans. After evaluation of the CS, as shown Fig. 13, the mineral admixtures can be classified as low reactive pozzolanic materials. The suspensions 2:8 for CS\_PN and CS\_PQ reached unsaturation with a loss of conductivity of 30% at 74 h and 156 h, respectively.

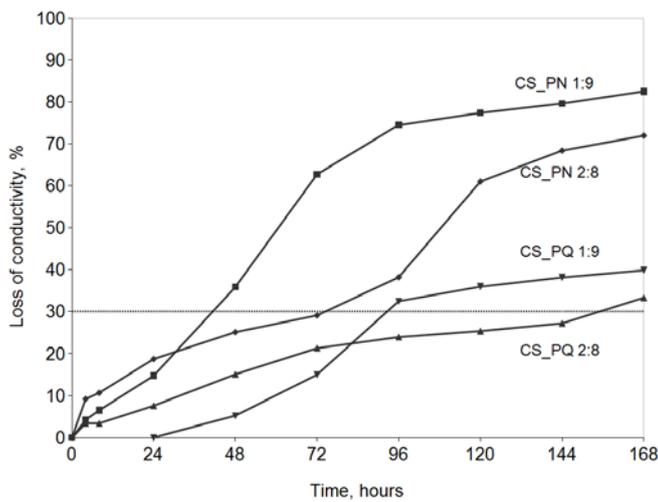


Fig. 12. Loss of electrical conductivity for CH: CS, 1:9 and 2:8 suspensions, test temperature of 60°C.

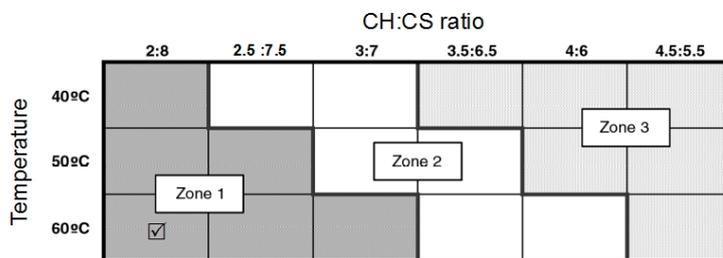


Fig. 13. Template for classification of pozzolan reactivity (Adapted from Tashima et al., 2014).

### 4. Conclusions

Both ancient CS (CS\_PQ and CS\_PN, in Chile) studied by means of the electrical conductimetric method for pozzolanic reactivity assessment can be classified as low reactive

pozzolanic materials. Nevertheless, the strength activity ratio (SAR) in PC mortars in which 25% of the binder is CS results agree with the requirements of the European standard EN 450-1 for fly ashes. The highest reactivity of CS\_PQ at 90 days agrees with the higher amorphous character of the sample shown in the X-ray diffraction pattern. CS AA pastes showed a DTG peak between 129 and 148°C, which is typical for AA gels. In general, an increase in the total weight loss for AA pastes occurs when curing times and curing temperature increase, providing evidence of the progression of the AA reaction. Again, reaction progress of AA for CS\_PQ was slightly higher than for CS\_PN. These results are in agreement with those obtained for the compressive strength of AA mortars, wherein the compressive strength of CS\_PQ mortars was substantially higher than those reported with the CS\_PN. Probably, the greatest CS\_PQ reactivity is related to the higher amorphous phase content, which facilitates the dissolution of the silica and alumina in the alkaline environment, favouring reaction with the alkaline solution. In accordance with the results, we can finally conclude that the studied ancient CS could be used successfully in PC and AA matrices, contributing to solving the environmental problem of the disposal of this waste, which has accumulated in large amounts in Chilean landfills.

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